

# Reviews

## *P*- $\rho$ -*T* Data of Liquids: Summarization and Evaluation. 6. Nonaromatic Hydrocarbons ( $C_n$ , $n \geq 5$ ) except *n*-Alkanes $C_5$ to $C_{16}$

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The published experimental data for 56 hydrocarbons (22 alkanes, 11 alkenes, 20 cycloalkanes, and 3 cycloalkenes) are summarized and reviewed, and the parameters of the Tait equation are given. This equation allows the calculation of smoothed values of either the volume ratio  $V(P)/V(P_{\text{ref}})$  and related properties (relative density  $\rho(P)/\rho(P_{\text{ref}})$  and compression  $\{1 - \rho(P_{\text{ref}})/\rho(P)\}$ ) or, using density data at atmospheric pressure ( $P_{\text{ref}} = 0.1$  MPa) or at saturation ( $P_{\text{ref}} = P_{\text{sat}}$ ), the liquid density of the substances over a temperature and pressure range. A comparison of isothermal compressibilities calculated from the Tait equation with available data at  $P = 0.1$  MPa from the literature is also presented.

### Introduction

This work is the last part of a systematic summarization and critical evaluation of published  $P$ - $\rho$ - $T$  data of hydrocarbons  $C_n$ ,  $n \geq 5$ , in a liquid state. The data for *n*-alkanes from  $C_5$  to  $C_{16}$  [96-cib/hne] and aromatic hydrocarbons [99-cib/tak] have already been reviewed and evaluated. This work concerns the remaining groups of hydrocarbons for which the available published experimental values of density,  $\rho(T, P)$  relative density  $\rho(T, P)/\rho(T, P = 0.1 \text{ MPa or } P_{\text{sat}})$ , and related quantities of liquids were compiled from the literature and evaluated. The present work and previous reviews [96-cib/hne, 99-cib/tak] do not include compilation or data evaluation for hydrocarbons with less than five carbon atoms; comprehensive reviews of thermodynamic properties (including  $P$ - $\rho$ - $T$  surface of the fluid state) for most of these hydrocarbons may be found in the literature [75-zor/hen, 81-bue/mau, 83-wax/gal, 86-jah/jac, 87-you/ely, 89-fri/ely, 91-fri/ing, 91-set/wag].

### Sources of Data

The original experimental data (9576 data points for 56 substances) processed were extracted from the source database which was employed for our previous reviews and is currently being updated. A list of substances is presented in Table 1 along with Chemical Abstracts Service Registry Numbers (CASRN) and formulas.

The characteristics of data that were available in the database for density and related quantities (molar and specific volumes, volume or density ratios, compression) of selected compounds are summarized in Table 2. The

temperature ranges and numbers of experimental values are restricted up to the critical temperature; that is, only subcritical liquid density data were taken from the source database. No corrections for the different temperature scales were made; the effect is mostly less than the uncertainties in density and/or temperature measurements. Very few researchers declare a particular temperature scale used (see "data type" column and footnotes f through i in Table 2). Similarly, as in our previous reviews, values (denoted by the letter F in the "data type" column of Table 2) calculated from smoothing functions presented in the papers (mostly the Tait equation), following as much as possible the information concerning the distribution of experimental points given by authors, were included for some substances if no direct experimental (D) or smoothed values (S) were available in the papers. The  $\rho(T, P)$  values calculated from other properties (C) were also included in the evaluation.

### Treatment of Data and Method of Data Evaluation

The procedures of treatment of the data and the critical evaluation were essentially the same as those employed in our previous papers [96-cib/hne, 99-cib/tak]. A brief summarization only is given below.

Available data on the compressed liquid density and related quantities were fitted by a Tait equation with the temperature-dependent parameters  $C(T)$  and  $B(T)$ , written in the form

$$\rho(T, P, \bar{v}, \bar{b}) = \frac{\rho(T, P_{\text{ref}}(T))}{1 - C(T, \bar{v}) \ln \left[ \frac{B(T, \bar{b}) + P}{B(T, \bar{b}) + P_{\text{ref}}(T)} \right]} \quad (1)$$

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**Table 1. List of Substances: Names (Alternative Names), Chemical Abstracts Service Registry Numbers (CASRN, supplied by authors), and Summary Formulas**

name (alternative name)	CASRN	formula	name (alternative name)	CASRN	formula
Alkanes					
2-methylbutane (isopentane)	78-78-4	C <sub>5</sub> H <sub>12</sub>	octadecane	593-45-3	C <sub>18</sub> H <sub>38</sub>
2,2-dimethylpropane (neopentane)	463-82-1	C <sub>5</sub> H <sub>12</sub>	7-hexyltridecane	7225-66-3	C <sub>19</sub> H <sub>40</sub>
2-methylpentane	107-83-5	C <sub>6</sub> H <sub>14</sub>	eicosane	112-95-8	C <sub>20</sub> H <sub>42</sub>
3-methylpentane	96-14-0	C <sub>6</sub> H <sub>14</sub>	docosane	629-97-0	C <sub>22</sub> H <sub>46</sub>
2,2-dimethylbutane	75-83-2	C <sub>6</sub> H <sub>14</sub>	tetracosane	646-31-1	C <sub>24</sub> H <sub>50</sub>
2,3-dimethylbutane	79-29-8	C <sub>6</sub> H <sub>14</sub>	9-octylheptadecane	7225-64-1	C <sub>25</sub> H <sub>52</sub>
3-ethylpentane	617-78-7	C <sub>7</sub> H <sub>16</sub>	triacontane	638-68-6	C <sub>30</sub> H <sub>62</sub>
2,2-dimethylpentane	590-35-2	C <sub>7</sub> H <sub>16</sub>	2,6,10,15,19,23-hexamethyltetracosane (squalane)	111-01-3	C <sub>30</sub> H <sub>62</sub>
2,2,3-trimethylbutane	464-06-2	C <sub>7</sub> H <sub>16</sub>	1,1-decylheneicosane	55320-06-4	C <sub>31</sub> H <sub>64</sub>
2,2,4-trimethylpentane (isooctane)	540-84-1	C <sub>8</sub> H <sub>18</sub>	13-dodecylhexacosane	55517-73-2	C <sub>38</sub> H <sub>78</sub>
heptadecane	629-78-7	C <sub>17</sub> H <sub>36</sub>	tetracontane	4181-95-7	C <sub>40</sub> H <sub>82</sub>
Alkenes					
2-methyl-1,3-butadiene (isoprene)	78-79-5	C <sub>5</sub> H <sub>8</sub>	1-hexene	592-41-6	C <sub>6</sub> H <sub>12</sub>
1-pentene	109-67-1	C <sub>5</sub> H <sub>10</sub>	1-heptene	592-76-7	C <sub>7</sub> H <sub>14</sub>
2-methyl-2-butene (amylene)	513-35-9	C <sub>5</sub> H <sub>10</sub>	1-octene (caprylene)	111-66-0	C <sub>8</sub> H <sub>16</sub>
(E)-1,4-hexadiene	7319-00-8	C <sub>6</sub> H <sub>10</sub>	1-nonene	124-11-8	C <sub>9</sub> H <sub>18</sub>
(Z,Z,E)-2,4-hexadiene (2-cis-4-trans-hexadiene)	5194-50-3	C <sub>6</sub> H <sub>10</sub>	2,6,10,15,19,23-hexamethyl-2,6,10,14,18,22-tetracosahexaene (squalene)	7683-64-9	C <sub>30</sub> H <sub>50</sub>
1,5-hexadiene	592-42-7	C <sub>6</sub> H <sub>10</sub>			
Cycloalkanes					
cyclopentane	287-92-3	C <sub>5</sub> H <sub>10</sub>	butylcyclohexane	1678-93-9	C <sub>10</sub> H <sub>20</sub>
methylcyclopentane	96-37-7	C <sub>6</sub> H <sub>12</sub>	bicyclohexyl	92-51-3	C <sub>12</sub> H <sub>22</sub>
cyclohexane	110-82-7	C <sub>6</sub> H <sub>12</sub>	octadecahydrochrysene (perhydrochrysene)	2090-14-4	C <sub>18</sub> H <sub>30</sub>
cyclohexane-d <sub>12</sub> (dodecadeuteriocyclohexane)	1735-17-7	C <sub>6</sub> D <sub>12</sub>	1,7-dicyclopentyl-4-(3-cyclopentylpropyl)heptane	55429-35-1	C <sub>25</sub> H <sub>46</sub>
methylcyclohexane	108-87-2	C <sub>7</sub> H <sub>14</sub>	1-(1-decahydronaphthyl)pentadecane	66359-82-8	C <sub>25</sub> H <sub>48</sub>
cycloheptane	291-64-5	C <sub>7</sub> H <sub>14</sub>	1-cyclopentyl-4-(3-cyclopentylpropyl)dodecane	7225-68-5	C <sub>25</sub> H <sub>48</sub>
ethylcyclohexane	1678-91-7	C <sub>8</sub> H <sub>16</sub>	1-cyclohexyl-3-(2-cyclohexylethyl)undecane	7225-69-6	C <sub>25</sub> H <sub>48</sub>
1-cis-2-dimethylcyclohexane	2207-01-4	C <sub>8</sub> H <sub>16</sub>	9-(2-cyclohexylethyl)heptadecane	25446-35-9	C <sub>25</sub> H <sub>50</sub>
cyclooctane	292-64-8	C <sub>8</sub> H <sub>16</sub>	9-(3-cyclopentylpropyl)heptadecane	5638-09-5	C <sub>25</sub> H <sub>50</sub>
trans-bicyclo[4.4.0]decane (trans-decalin)	493-02-7	C <sub>10</sub> H <sub>18</sub>	1,1-bis(decacyclo-1-naphthyl)undecane	55373-96-1	C <sub>31</sub> H <sub>56</sub>
Cycloalkenes					
1,3-cyclohexadiene	592-57-4	C <sub>6</sub> H <sub>8</sub>	cyclohexene	110-83-8	C <sub>6</sub> H <sub>10</sub>
1,4-cyclohexadiene	628-41-1	C <sub>6</sub> H <sub>8</sub>			

where

$$C(T, \bar{c}) = \sum_{i=0}^{N_C} c_i [(T - T_0)/100]^i$$

$$\bar{c} = \{c_i\} = \{c_0, \dots, c_{N_C}\} \quad (2)$$

$$B(T, \bar{b}) = \sum_{i=0}^{N_B} b_i [(T - T_0)/100]^i$$

$$\bar{b} = \{b_i\} = \{b_0, \dots, b_{N_B}\} \quad (3)$$

and  $T_0$  is a parameter with a preselected fixed value for which  $C(T_0) = c_0$  and  $B(T_0) = b_0$  are valid. The reference values  $\rho(T, P_{\text{ref}}(T))$  and  $P_{\text{ref}}(T)$  were selected in the same way as that used previously; that is, at temperatures below the normal boiling temperature, the densities at atmospheric pressure ( $P_{\text{ref}} = 0.101\,325$  MPa) were used, while for higher temperatures the values along the saturation curve, that is, saturated liquid densities and saturated vapor pressures, were employed. Experimental values of densities at atmospheric pressure or at saturation for the same sample reported along with compressed liquid density data were preferably used for the reference density  $\rho(T, P_{\text{ref}})$ , and thus the values of relative density  $\rho(T, P)/\rho(T, P_{\text{ref}} = 0.1$  MPa or  $P_{\text{sat}})$  reported by the authors were correlated by eq 1. In some cases of isothermal data, the reference density  $\rho(T, P_{\text{ref}})$  was obtained for each isotherm by an extrapolation of experimental compressed liquid density data to the reference pressure  $P_{\text{ref}}$  (0.101 325 MPa below or  $P_{\text{sat}}$  above the normal boiling temperature), using the Tait equation. If the reference values were not available in the original source, and the extrapolation was not feasible (e.g., for

isobaric or isochoric data), then the densities obtained from the equations summarized in the Appendix (Table 6) were employed in the correlations. Saturated vapor pressures were calculated from the smoothing functions taken from the literature (for references, see Table 3) and used in the correlations.

The adjustable parameters  $\bar{c}$  and  $\bar{b}$  of function 1 were obtained by minimizing the objective function

$$\phi(\bar{c}, \bar{b}) = \sum_{j=1}^{N_p} w_j [\rho_j - \rho(T_j, P_j, \bar{c}, \bar{b})]^2 \quad (4)$$

where  $\rho_j$ ,  $T_j$ ,  $P_j$  is the  $j$ th experimental data point,  $\rho(T_j, P_j, \bar{c}, \bar{b})$  is the value calculated from function 1 with the parameters  $\bar{c}$  and  $\bar{b}$  for the values  $T_j$  and  $P_j$ , and  $N_p$  is the number of experimental values of density used in the correlation. The adjustable parameters were calculated by the Marquardt algorithm in double precision to minimize the influence of rounding errors. The statistical weights  $w_j$  in eq 4 were defined as

$$w_j = \mu_j / (\delta\rho_j)^2 \quad (5)$$

where  $\delta\rho_j$  is the experimental uncertainty taken from the source database and either given by the authors (preferably) or estimated by a compiler for the  $j$ th density value in a correlated data set. The uncertainties  $\delta\rho_j$  included not only random but also systematic error estimates (if available) and corresponded to the experimental accuracy rather than the precision of measurements. The statistical weight of each density value was adjusted by varying the parameter  $\mu_j$  ( $\mu_j = 0$  for rejected values), taking into account

**Table 2. Characteristics of Data Sets: Overall Number of Data Points ( $N_p$ ), Temperature and Pressure Ranges within the Liquid State ( $T_{\min}$ ,  $T_{\max}$ ,  $P_{\min}$ , and  $P_{\max}$ ), Experimental Method Used, Types of Data, and Purities of Measured Samples**

ref	$N_p$	$T_{\min}/K$	$T_{\max}/K$	$P_{\min}/MPa$	$P_{\max}/MPa$	method <sup>a</sup>	data type <sup>b</sup>	sample purity <sup>c</sup> /%
2-Methylbutane								
31-bri	29	273.15	368.15	49.1	882.7	vb	D	
54-isa/li	34	373.15	448.15	1.1	21.6	rl	D	99.89m% <sup>e</sup>
69-mop	37	223.15	298.15	10.2	202.8	vb	D	99.99 <sup>e</sup>
71-hou/hey	15	298.15	298.15	50.0	2400.0	vs	S	
74-hou	12	295.15	295.15	200.0	2400.0	vs	D	
76-sah/gag	7	293.15	293.15	1.0	7.0	ce	F	99.9m <sup>e</sup>
92-wal/bar	180	201.40	259.90	10.0	300.0	vs	S	99.5 <sup>e</sup>
total	314	201.40	448.15	1.0	2400.0			
2,2-Dimethylpropane								
68-gon/lee	18	310.93	410.93	0.7	55.2	hp	D	99.92m <sup>e</sup>
73-daw/sil	61	343.15	393.15	0.6	31.6	vl	D	>99.9 <sup>e</sup>
73-koh/luk	38	298.15	298.15	0.3	6.9	vl	F	>99 <sup>d</sup>
75-luk/dav	49	298.15	298.15	0.3	6.6	vl	F	>99m <sup>d</sup>
total	166	298.15	410.93	0.3	55.2			
2-Methylpentane								
31-bri	34	273.15	368.15	49.1	1176.9	vb	D	
40-kel/fel	37	373.15	473.15	0.6	31.6	pi	D <sup>f</sup>	
76-sah/gag	14	293.15	298.15	1.0	7.0	ce	F	99.9m <sup>e</sup>
88-mor/aon	64	298.15	313.15	1.8	135.8	vl	D	
total	149	273.15	473.15	0.6	1176.9			
3-Methylpentane								
31-bri	40	273.15	368.15	49.1	1176.9	vb	D	
52-day/fel	76	353.15	498.15	0.6	31.6	vl	D	99.80m <sup>e</sup>
76-sah/gag	14	293.15	298.15	1.0	7.0	ce	F	99.9m <sup>d</sup>
88-mor/aon	63	298.15	313.15	2.5	145.4	vl	D	
total	193	273.15	498.15	0.6	1176.9			
2,2-Dimethylbutane								
31-bri	26	273.15	368.15	49.1	980.8	vb	D	
43-fel/wat	69	373.15	473.15	1.0	30.4	vl	S <sup>f</sup>	
76-sah/gag	14	293.15	298.15	1.0	7.0	ce	F	99.9m <sup>e</sup>
88-mor/aon	58	298.15	313.15	1.5	124.3	vl	D	
90-pol/wei	72	243.70	313.20	5.0	200.0	vb	F	99.0 <sup>d</sup>
total	239	243.70	473.15	1.0	980.8			
2,3-Dimethylbutane								
31-bri	32	273.15	368.15	49.1	1078.8	vb	D	
42-kel/fel	44	373.15	498.15	0.6	31.6	vl	D <sup>f</sup>	
76-sah/gag	14	293.15	298.15	1.0	7.0	ce	F	99.9m <sup>d</sup>
87-hol/goe	5	293.15	293.15	2.0	10.0	mo	D <sup>g</sup>	>99.6m <sup>e</sup>
88-mor/aon	61	298.15	313.15	1.6	139.3	vl	D	
90-rie/sch	365	173.20	303.20	10.0	300.0	vs	S	99.5 <sup>d</sup>
93-bao/cac	215	208.16	298.15	0.2	108.5	rl	D <sup>h</sup>	99.2m <sup>e</sup>
total	736	173.20	498.15	0.2	1078.8			
3-Ethylpentane								
70-kus/tas	20	298.15	353.15	39.2	196.1	vl	D	
2,2-Dimethylpentane								
76-sah/gag	14	293.15	298.15	1.0	7.0	ce	F	99.9m <sup>e</sup>
2,2,3-Trimethylbutane								
70-kus/tas	11	298.15	353.15	39.2	156.9	vl	D	
2,2,4-Trimethylpentane								
32-bri	30	273.15	368.15	49.0	980.7	vb	D	
43-fel/wat-1	72	373.15	523.15	0.5	30.4	vl	S	99.9m <sup>e</sup>
85-dym/isd	43	298.19	372.95	25.4	539.8	vb	D <sup>g</sup>	>99m <sup>d</sup>
86-hol/goe	5	293.15	293.15	2.0	10.0	mo	D <sup>g</sup>	>99.8m <sup>e</sup>
87-led	70	298.15	358.15	10.0	330.0	vb	S	99.5 <sup>d</sup>
90-mal/woo	115	278.15	323.15	2.5	280.0	vb	F	99.8m <sup>d</sup>
90-mal/woo	46	323.15	338.15	2.5	280.0	vb	F	99.8m <sup>d</sup>
90-pol/wei	48	273.15	348.15	5.0	200.0	vb	F	99.5 <sup>d</sup>
91-pap/zia	9	298.15	298.15	2.0	33.8	mo	D	99.5 <sup>d</sup>
92-naz/gas	77	294.15	538.65	5.0	58.9	bu	D	>99.98 <sup>d</sup>
93-mal/woo	89	313.15	353.15	2.6	374.3	vb	D	
94-pad/far	30	298.15	348.15	10.1	100.1	mo/bu	D	99.8 <sup>e</sup>
96-hah/ulc	5	293.15	293.15	2.0	10.0	mo	D	99.92m <sup>e</sup>
96-pad/far-1	47	197.93	298.15	10.1	100.1	mo/bu	D <sup>h</sup>	99.7 <sup>e</sup>
96-pad/far-1	30	298.15	348.16	10.1	100.1	mo/bu	D <sup>h</sup>	99.7 <sup>e</sup>
total	716	197.93	538.65	0.5	980.7			
Heptadecane								
64-doo	60	323.15	573.15	5.0	500.0	vl	S	
87-man/cri	27	298.15	338.15	2.0	10.1	vl	F	99m <sup>e</sup>
total	87	298.15	573.15	2.0	500.0			

Table 2. Continued

ref	$N_p$	$T_{min}/K$	$T_{max}/K$	$P_{min}/MPa$	$P_{max}/MPa$	method <sup>a</sup>	data type <sup>b</sup>	sample purity <sup>c</sup> /%
Octadecane								
58-cut/mcm	48	333.15	408.15	34.5	551.3	vb	D	
7-Hexyltridecane								
59-low/spe	38	310.93	408.15	20.0	340.0	vb	S	
Eicosane								
64-doo	50	373.15	573.15	5.0	500.0	vl	S	
Docosane								
88-pet/spi	48	323.08	368.26	2.1	16.1	ct	D	>99m <sup>e</sup>
Tetracosane								
87-pet/van	36	333.52	371.22	2.1	12.1	nd	D	>98m <sup>e</sup>
9-Octylheptadecane								
58-cut/mcm	140	310.95	408.15	34.5	1033.7	vb	S	
59-low/spe	20	310.93	372.04	20.0	340.0	vb	S	
total	160	310.93	408.15	20.0	1033.7			
Triacontane								
64-doo	50	373.15	573.15	5.0	500.0	vl	S	
2,6,10,15,19,23-Hexamethyltetracosane								
70-kus/tas	20	298.15	353.15	39.2	196.1	vl	D	
11-Decylheneicosane								
59-low/spe	36	310.93	408.15	20.0	340.0	vb	S	
13-Dodecylhexacosane								
59-low/spe	34	310.93	408.15	20.0	340.0	vb	S	
Tetracontane								
64-doo	40	423.15	573.15	5.0	500.0	vl	S	
2-Methyl-1,3-butadiene								
32-bri	14	273.15	273.15	49.0	1176.8	vb	D	
1-Pentene								
51-day/fel	50	353.15	448.15	0.6	31.6	vl	D	99.34m <sup>e</sup>
2-Methyl-2-butene								
49-bri	10	298.15	298.15	49.0	490.3	vs	D	
(E)-1,4-Hexadiene								
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	99m <sup>d</sup>
(2Z,4E)-2,4-Hexadiene								
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	99m <sup>d</sup>
1,5-Hexadiene								
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	99.9m <sup>d</sup>
1-Hexene								
70-aba/ker	192	283.15	503.15	4.0	68.7	bu	D	99.8 <sup>e</sup>
72-ker/apa	431	283.15	503.15	0.8	68.7	bu	D	99.98 <sup>e</sup>
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	99.9m <sup>d</sup>
82-gus/gal	48	146.00	293.50	5.0	50.0	bu	D	
total	681	146.00	503.15	0.8	68.7			
1-Heptene								
81-gus/naz	78	293.15	523.15	4.0	50.0	bu	D	
1-Octene								
49-bri	17	298.15	298.15	49.0	3922.7	vs	D	
72-ker/apa	361	283.15	533.15	2.6	68.7	bu	D	99.98 <sup>e</sup>
88-dym/mal	33	298.15	373.15	10.6	312.7	vb	D	99m <sup>d</sup>
92-naz/gas	70	290.15	538.95	5.0	58.9	bu	D	>99.98 <sup>d</sup>
total	481	283.15	538.95	2.6	3922.7			
1-Nonene								
84-gus/gal	70	198.00	523.00	10.0	50.0	bu	D	
2,6,10,15,19,23-Hexamethyl-2,6,10,14,18,22-tetracosahexaene								
70-kus/tas	20	298.15	353.15	39.2	196.1	vl	D	
Cyclopentane								
69-bra/fre	9	303.15	303.15	50.0	450.0	vb	D	
70-kus/tas	20	298.15	353.15	39.2	196.1	vl	D	
92-bao/cac	186	192.79	298.15	0.1	104.3	rl	D <sup>b</sup>	>99.5m <sup>d</sup>
total	215	192.79	353.15	0.1	450.0			
Methylcyclopentane								
80-oza/ooy	41	298.20	348.20	9.9	196.2	vl	D	99.9 m <sup>e</sup>
87-hol/goe	5	293.15	293.15	2.0	10.0	mo	D <sup>g</sup>	>99.6m <sup>e</sup>
total	46	293.15	348.20	2.0	196.2			

Table 2. Continued

ref	$N_p$	$T_{min}/K$	$T_{max}/K$	$P_{min}/MPa$	$P_{max}/MPa$	method <sup>a</sup>	data type <sup>b</sup>	sample purity <sup>c</sup> /%
Cyclohexane								
34-rot/nag	9	289.29	291.47	0.2	20.9	vl	D	
40-sch/hof	3	311.08	377.59	5.5	5.5	vl	D	
57-rea/sag	78	310.93	510.93	0.4	68.5	vl	D <sup>i</sup>	>99.8m <sup>e</sup>
59-gol/vag	89	293.65	548.15	0.3	53.5	bu	D	
59-gol/vag	80	293.15	553.15	5.1	50.7	bu	S	
62-hol/wha	50	298.16	348.15	1.0	10.0	vl	F	>99.9 <sup>e</sup>
70-kus/tas	6	313.15	353.15	39.2	117.7	vl	D	
72-gol/ada	30	303.15	393.15	50.7	253.3	vb	D	
72-ker/apa	541	283.15	553.15	1.1	68.7	bu	D	99.98 <sup>e</sup>
72-lys	2	297.15	297.15	179.0	421.0	sw	D	
73-ker/apa	368	293.15	553.16	1.1	68.7	bu	D	99.98 <sup>e</sup>
73-rog/bur	10	298.15	298.15	1.0	10.0	ce	F	
74-apa/ker	101	283.15	553.15	2.1	68.7	bu	D	>99.98 <sup>e</sup>
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	
75-gri/mur	40	498.15	548.15	2.1	79.1	pi	D	99.90 <sup>e</sup>
75-ras/gri	96	298.15	473.15	0.6	85.4	pi	D	99.90 <sup>e</sup>
75-ras/gri	35	298.15	473.15	1.6	71.0	pi	D	99.90 <sup>e</sup>
78-gou	75	314.15	393.15	5.1	40.1	vl	D	
79-dic	18	295.00	295.00	1800.0	42600.0	sw	D	99 <sup>d</sup>
79-isd/dym	4	298.15	348.15	10.0	20.0	ul	C	>99 <sup>e</sup>
79-isd/dym	2	348.15	348.15	50.0	100.0	vb	D	>99 <sup>e</sup>
79-kas/fuk	32	298.15	348.15	6.9	105.3	vl	D	99.8v <sup>d</sup>
80-jon/has	40	313.00	383.00	5.0	214.0	vb	D	
82-wis/wue	27	286.80	337.90	10.0	110.0	vs	S	99.9 <sup>e</sup>
84-mat/van	34	288.15	313.15	2.5	35.0	mo	F	
87-sun/kor	80	288.15	323.06	5.0	85.0	ul	C <sup>g</sup>	>99.5v <sup>d</sup>
88-mor/aon	27	298.15	313.15	2.4	46.1	vl	D	
89-vos/slo	24	326.50	450.00	0.4	10.4	vs	D	
90-pol/wei	35	293.20	368.10	5.0	180.0	vb	F	99.9 <sup>d</sup>
90-tos/fig	48	308.07	343.15	4.6	101.9	bu	D	99.5m <sup>d</sup>
90-tos/fig	6	333.15	333.15	39.8	91.7	bu	D	99.5m <sup>d</sup>
91-mel/mel	49	293.15	413.15	10.0	250.0	pi	D	
91-tan/hos	23	298.15	348.15	6.2	100.0	va	D	99.8 <sup>e</sup>
92-lag/bon	18	288.15	323.15	5.0	15.0	mo <sup>j</sup>	D	
92-lag/bon	18	288.15	323.15	5.0	15.0	mo <sup>j</sup>	D	
96-pad/far	20	298.15	348.14	5.1	62.8	mo/bu	D	99.85 <sup>e</sup>
total	2128	283.15	553.16	0.2	42600.0			
Cyclohexane- <i>d</i> <sub>12</sub>								
84-mat/van	35	288.15	313.15	2.5	35.0	mo	F	
Methylcyclohexane								
49-bri	11	298.15	298.15	49.0	980.7	vs	D	
69-bra/fre	9	303.15	303.15	50.0	450.0	vb	D	
70-aba/ker	255	283.15	568.15	5.0	68.7	bu	D	99.9 <sup>e</sup>
72-ker/apa	459	283.15	568.15	5.0	68.7	bu	D	99.98 <sup>e</sup>
72-ker/apa	38	283.15	568.15	0.3	6.0	bu	D	99.98 <sup>e</sup>
78-gou	90	293.15	392.55	5.1	40.1	vl	D	
79-jon/has	27	203.00	298.00	50.0	500.0	nd	D	
97-bay/bon	15	303.15	343.15	20.0	100.0	mo	D	>99 <sup>d</sup>
total	904	203.00	568.15	0.3	980.7			
Cycloheptane								
70-kus/tas	12	298.15	353.15	39.2	196.1	vl	D	
78-gou	90	293.65	393.15	5.1	40.1	vl	D	
total	102	293.65	393.15	5.1	196.1			
Ethylcyclohexane								
83-gus/sha	78	293.15	523.15	5.0	50.0	bu	D	
89-vos/slo	24	327.10	440.90	0.4	9.7	vs	D	
total	102	293.15	523.15	0.4	50.0			
1- <i>cis</i> -2-Dimethylcyclohexane								
89-vos/slo	21	325.20	461.20	0.4	9.7	vs	D	
Cyclooctane								
78-gou	75	313.65	393.85	5.1	40.1	vl	D	
<i>trans</i> -Bicyclo[4.4.0]decane ( <i>trans</i> -decalin)								
70-kus/tas	20	298.15	353.15	39.2	196.1	vl	D	
Butylcyclohexane								
87-hol/goe	5	293.15	293.15	2.0	10.0	mo	D <sup>g</sup>	>99.8m <sup>e</sup>
Bicyclohexyl								
88-sid/tej	21	298.20	338.20	0.7	34.5	mo	D	>99m <sup>d</sup>

Table 2. Continued

ref	$N_p$	$T_{\min}/K$	$T_{\max}/K$	$P_{\min}/MPa$	$P_{\max}/MPa$	method <sup>a</sup>	data type <sup>b</sup>	sample purity <sup>c</sup> /%
Octadecahydrochrysene								
59-low/spe	48	310.93	408.15	20.0	340.0	vb	S	
1,7-Dicyclopentyl-4-(3-cyclopentylpropyl)heptane								
58-cut/mcm	113	310.95	408.15	34.5	1033.7	vb	S	
1-(1-Decahydronaphthyl)pentadecane								
58-cut/mcm	49	333.15	408.15	34.5	585.7	vb	S	
1-Cyclopentyl-4-(3-cyclopentylpropyl)dodecane								
58-cut/mcm	149	310.95	408.15	34.5	1033.7	vb	S	
1-Cyclohexyl-3-(2-cyclohexylethyl)undecane								
58-cut/mcm	103	310.95	408.15	34.5	895.8	vb	S	
9-(2-Cyclohexylethyl)heptadecane								
58-cut/mcm	126	310.95	408.15	34.5	1033.7	vb	S	
9-(3-Cyclopentylpropyl)heptadecane								
58-cut/mcm	139	310.95	408.15	34.5	1033.7	vb	S	
1,1-Bis(decahydro-1-naphthyl)undecane								
59-low/spe	35	310.93	408.15	20.0	300.0	vb	S	
1,3-Cyclohexadiene								
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	99m <sup>d</sup>
79-dic	11	295.00	295.00	920.0	41100.0	sw	D	97 <sup>d</sup>
total	21	293.15	298.15	1.0	41100.0			
1,4-Cyclohexadiene								
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	99.7m <sup>d</sup>
79-dic	11	295.00	295.00	950.0	41700.0	sw	D	97 <sup>d</sup>
total	21	293.15	298.15	1.0	41700.0			
Cyclohexene								
72-ker/apa	364	283.15	533.15	2.6	68.7	bu	D	99.98 <sup>e</sup>
75-bur/ric	10	293.15	298.15	1.0	5.0	ce	F	99.9m <sup>d</sup>
79-dic	15	295.00	295.00	500.0	40300.0	sw	D	99 <sup>d</sup>
91-mel/mel	70	293.15	413.15	10.0	250.0	pi	D	
total	459	283.15	533.15	1.0	40300.0			

<sup>a</sup> Method used for measurements: bu, buoyancy method; ca, densities obtained by integration from thermal expansivities measured by calorimetric method; ce, densities evaluated by integration from isothermal compressibilities obtained by ultracentrifuge method; ct, Cailletet apparatus; hp, high-pressure pycnometer; mo, mechanical oscillator method; mo/bu, combination of vibrating-wire technique with buoyancy principle; nd, not described or stated in the reference; pi, piezometer of unspecified type; rl, expansion principle; sw, shock wave method; ul, densities evaluated from speeds of ultrasound; va, Aime method; vb, variable-volume cell with bellows; vl, variable-volume cell with liquid piston; vs, variable-volume cell with solid piston. For the classification and description of the methods, see ref [85-tek/cib]. <sup>b</sup> D, direct experimental data; S, smoothed data presented in the reference; C, calculated from other properties; F, values calculated from smoothing equation reported by the researchers. <sup>c</sup> No letter, unspecified percent; m, mole percent; v, volume percent; w, mass percent. <sup>d</sup> Purity of source material is given only. <sup>e</sup> Final purity of the sample. <sup>f</sup> ITS-27 declared by the researchers. <sup>g</sup> IPTS-68 declared by the researchers. <sup>h</sup> ITS-90 declared by the researchers. <sup>i</sup> ITS-48 declared by the researchers. <sup>j</sup> Two methods of calibration of DMA 512 vibrating-tube densimeter reported; the first one declared to be superior.

additional available information (sample purity, experimental method used, uncertainties in temperature and pressure measurements). In some cases comparisons of isothermal compressibilities calculated from the fit of a particular data set with independent values (see below and Table 5) were made to facilitate the adjustment. The calculations of the parameters  $\bar{c}$  and  $\bar{b}$  were repeated until the final fit was obtained where the deviations between retained experimental and smoothed values were roughly equal to the modified experimental uncertainties,  $\delta\rho_j\mu_j^{1/2}$ , that is, where the weighted standard deviation of the fit was close to unity. In those cases where only one set of smoothed values (S-type data, Table 2) was available for a particular substance, the statistical weights were not modified ( $\mu_j = 1$  for retained and  $\mu_j = 0$  for rejected data points) and therefore the weighted standard deviation  $s_w$  (see below) may differ from unity.

## Results

Table 3 records the values of the parameters of eq 1 for each substance along with some statistical information of

the fits defined as follows:

$$\text{RMSD} = \left\{ \sum_{j=1}^{N_p} [\rho_j - \rho(T_j, P_j, \bar{c}, \bar{b})]^2 / N_p \right\}^{1/2} \quad (6)$$

$$\text{RMSD}_r/\% = 100 \left\{ \sum_{j=1}^{N_p} [1 - \rho(T_j, P_j, \bar{c}, \bar{b}) / \rho_j]^2 / N_p \right\}^{1/2} \quad (7)$$

$$\text{bias} = \sum_{j=1}^{N_p} [\rho_j - \rho(T_j, P_j, \bar{c}, \bar{b})] / N_p \quad (8)$$

$$\pm = \sum_{j=1}^{N_p} \text{sign}[\rho_j - \rho(T_j, P_j, \bar{c}, \bar{b})] \cdot 1 \quad (9)$$

$$s_w = [\phi / (N_p - N_C - N_B - 2)]^{1/2} \quad (10)$$

where  $N_p$  is the overall number of experimental data points retained for the correlation. The characteristics are given on an absolute density scale ( $\text{kg}\cdot\text{m}^{-3}$ ), which is more illustrative than on a relative density scale.

**Table 3. Parameters  $c_0$ ,  $b_i$ , and  $T_0$  of Eq 1, Temperature and Pressure Ranges<sup>a</sup> ( $T_{\min}$ ,  $T_{\max}$ ,  $P_{\min}$ , and  $P_{\max}$ ), Absolute (RMSD) and Relative (RMSD<sub>r</sub>) Root Mean Square Deviations, Biases (Bias), Number of Data Points ( $N_p$ ),  $\pm$ , Weighted Standard Deviations ( $s_w$ ), and References to Saturated Vapor Pressure (ref( $P_{\text{sat}}$ )) for the Fits Where  $T_{\max}$  Is Higher Than the Normal Boiling Temperature**

	2-methyl-butane	2,2-dimethyl-propane	2-methyl-pentane	3-methyl-pentane	2,2-dimethyl-butane	2,3-dimethyl-butane	2,3-dimethyl-butane <sup>b</sup>	3-ethyl-pentane
$c_0$	0.088008	0.086510	0.089537	0.087915	0.088950	0.073172	0.087435	0.091852
$c_1/\text{K}^{-1}$						0.004085		
$b_0/\text{MPa}$	97.8313	27.3297	49.8393	48.9510	45.1214	39.8600	48.8186	67.8829
$b_1/(\text{MPa}\cdot\text{K}^{-1})$	-84.7915	-34.2103	-49.6496	-44.2980	-43.9956	-38.8467	-40.2085	-57.8087
$b_2/(\text{MPa}\cdot\text{K}^{-2})$	23.2045	7.6955	14.4427	9.2935	10.1425	11.2092	-30.9644	19.1434
$b_3/(\text{MPa}\cdot\text{K}^{-3})$			-1.3079			-1.2447		
$T_0/\text{K}$	201.40	298.15	298.15	303.15	298.15	298.15	298.15	298.15
$T_{\min}/\text{K}$	223.15	298.15	273.15	293.15	243.70	208.16	293.15	298.15
$T_{\max}/\text{K}$	298.15	393.15	473.15	473.15	473.15	473.15	313.15	353.15
$P_{\min}/\text{MPa}$	1.00	0.30	0.57	0.57	1.00	0.18	1.00	39.23
$P_{\max}/\text{MPa}$	2400.00	31.58	1176.90	145.40	200.00	106.25	139.30	196.13
RMSD/( $\text{kg}\cdot\text{m}^{-3}$ )	0.440	0.207	2.199	0.453	0.496	0.477	0.189	0.128
RMSD <sub>r</sub> /%	0.057	0.041	0.272	0.086	0.089	0.072	0.027	0.017
bias/( $\text{kg}\cdot\text{m}^{-3}$ )	-0.115	0.041	-0.042	0.040	-0.016	-0.165	-0.049	0.000
$N_p$	56	148	141	140	210	280	80	20
$\pm$	-10	-2	11	-8	20	-104	-20	0
$s_w$	0.914	1.092	1.140	1.140	1.095	1.214	1.057	1.012
ref( $P_{\text{sat}}$ )		83-mcg	83-mcg	83-mcg	83-mcg	83-mcg		

	2,2-dimethyl-pentane	2,2,3-trimethyl-butane	2,2,4-trimethyl-pentane	heptadecane	octadecane	7-hexyltridecane	eicosane	docosane
$c_0$	0.079479	0.090175	0.087786	0.087728	0.087801	0.088830	0.086743	0.039049
$b_0/\text{MPa}$	50.5418	37.2254	46.4381	89.5233	56.3108	68.8356	68.2027	38.0546
$b_1/(\text{MPa}\cdot\text{K}^{-1})$	-59.0336	-29.6553	-39.8361	-53.2648	-34.8999	-44.6628	-43.1426	-43.4898
$b_2/(\text{MPa}\cdot\text{K}^{-2})$		23.8796	10.6744	8.7057	14.5622	8.8493	7.6667	39.1505
$b_3/(\text{MPa}\cdot\text{K}^{-3})$			-2.4092					
$b_4/(\text{MPa}\cdot\text{K}^{-4})$			0.5894					
$T_0/\text{K}$	293.15	353.15	323.15	323.15	408.15	372.04	373.15	323.08
$T_{\min}/\text{K}$	293.15	298.15	197.93	298.15	333.15	310.93	373.15	323.08
$T_{\max}/\text{K}$	298.15	353.15	523.15	573.15	408.15	408.15	573.15	368.26
$P_{\min}/\text{MPa}$	1.00	39.23	0.51	2.03	34.46	20.00	5.00	2.05
$P_{\max}/\text{MPa}$	7.00	156.91	882.60	500.00	551.28	340.00	500.00	16.05
RMSD/( $\text{kg}\cdot\text{m}^{-3}$ )	0.005	0.082	0.639	0.944	0.403	0.230	1.222	0.422
RMSD <sub>r</sub> /%	0.001	0.011	0.091	0.121	0.051	0.028	0.159	0.055
bias/( $\text{kg}\cdot\text{m}^{-3}$ )	0.001	-0.001	-0.082	0.035	0.023	0.027	0.067	-0.001
$N_p$	14	11	446	87	48	36	50	48
$\pm$	2	1	-56	-11	-4	-2	0	-16
$s_w$	0.007	0.920	1.124	0.853	0.907	0.899	0.797	1.138
ref( $P_{\text{sat}}$ )			83-mcg					

	9-octyl-tetracosane	heptadecane	triacontane	HMTC <sup>c</sup>	11-decyl-heneicosane	13-dodecyl-hexacosane	tetracontane	2-methyl-1,3-butadiene
$c_0$	0.057048	0.092203	0.086183	0.090208	0.088556	0.089487	0.086751	0.099097
$b_0/\text{MPa}$	54.7838	84.8571	81.9415	122.1435	84.9378	90.1315	66.5183	80.8813
$b_1/(\text{MPa}\cdot\text{K}^{-1})$	-31.0645	-48.6706	-53.8673	-52.1931	-46.0832	-46.9989	-41.3867	
$b_2/(\text{MPa}\cdot\text{K}^{-2})$		12.2404	15.3258	-4.5813	6.0890	8.1517	9.0745	
$b_3/(\text{MPa}\cdot\text{K}^{-3})$			-2.0531					
$T_0/\text{K}$	333.52	372.05	373.15	298.15	372.04	372.04	423.15	273.15
$T_{\min}/\text{K}$	333.52	310.93	373.15	298.15	310.93	310.93	423.15	273.15
$T_{\max}/\text{K}$	371.22	408.15	573.15	353.15	408.15	408.15	573.15	273.15
$P_{\min}/\text{MPa}$	2.07	20.00	5.00	39.23	40.00	20.00	5.00	49.03
$P_{\max}/\text{MPa}$	12.07	1033.65	500.00	196.13	340.00	340.00	500.00	1176.80
RMSD/( $\text{kg}\cdot\text{m}^{-3}$ )	0.195	0.655	0.973	0.057	0.302	0.226	1.267	0.568
RMSD <sub>r</sub> /%	0.026	0.075	0.124	0.007	0.036	0.027	0.159	0.064
bias/( $\text{kg}\cdot\text{m}^{-3}$ )	0.000	0.060	0.099	0.000	0.022	0.029	0.142	0.041
$N_p$	36	136	47	20	31	30	40	14
$\pm$	-12	4	11	0	1	4	8	2
$s_w$	0.258	0.910	0.621	0.284	0.260	0.193	0.799	0.906

	1-pentene	2-methyl-2-butene	(E)-1,4-hexadiene	(2Z,4E)-2,4-hexadiene	1,5-hexadiene	1-hexene	1-heptene	1-octene
$c_0$	0.080544	0.096076	0.076611	0.075707	0.084083	0.085421	0.136085	0.089528
$b_0/\text{MPa}$	6.2555	51.9434	51.6052	60.1053	53.5627	51.1564	117.6648	67.6214
$b_1/(\text{MPa}\cdot\text{K}^{-1})$	-20.6345		-55.3701	-60.7677	-54.8755	-49.5730	-105.5210	-39.3409
$b_2/(\text{MPa}\cdot\text{K}^{-2})$	9.3081					14.4725		1.7555
$b_3/(\text{MPa}\cdot\text{K}^{-3})$						-1.0687		1.2891
$b_4/(\text{MPa}\cdot\text{K}^{-4})$						-0.2793		
$T_0/\text{K}$	398.15	298.15	293.15	293.15	293.15	298.15	293.15	298.15
$T_{\min}/\text{K}$	353.15	298.15	293.15	293.15	293.15	146.00	293.15	290.15
$T_{\max}/\text{K}$	448.15	298.15	298.15	298.15	298.15	473.15	353.15	538.95
$P_{\min}/\text{MPa}$	0.57	49.03	1.00	1.00	1.00	0.81	5.00	5.00
$P_{\max}/\text{MPa}$	31.60	490.33	5.00	5.00	5.00	68.75	50.00	312.70
RMSD/( $\text{kg}\cdot\text{m}^{-3}$ )	0.453	0.205	0.002	0.002	0.001	1.017	1.162	1.337
RMSD <sub>r</sub> /%	0.090	0.027	0.000	0.000	0.000	0.159	0.169	0.205
bias/( $\text{kg}\cdot\text{m}^{-3}$ )	-0.056	0.014	0.000	0.000	0.000	-0.246	-0.165	0.321
$N_p$	50	10	10	10	10	418	23	100
$\pm$	-20	-4	2	2	-2	-106	-7	14
$s_w$	1.066	1.150	0.036	0.029	0.013	1.153	1.200	1.659
ref( $P_{\text{sat}}$ )	83-mcg					83-mcg		79-dyk/rep

Table 3. Continued

	1-nonene	HMTE <sup>d</sup>	cyclopentane	methyl- cyclopentane	cyclohexane	cyclohexane- <i>d</i> <sub>12</sub>	methyl- cyclohexane
<i>c</i> <sub>0</sub>	0.177574	0.091245	0.088558	0.088019	0.085159	0.079666	0.086999
<i>b</i> <sub>0</sub> /MPa	381.1410	139.3681	67.5860	55.0075	60.3130	70.2463	80.3279
<i>b</i> <sub>1</sub> /(MPa·K <sup>-1</sup> )	-215.6507	-59.0280	-61.0179	-50.0741	-53.8926	-65.3112	-63.5774
<i>b</i> <sub>2</sub> /(MPa·K <sup>-2</sup> )	36.9232		17.2997	-8.2772	17.0338	-39.6447	17.2095
<i>b</i> <sub>3</sub> /(MPa·K <sup>-3</sup> )			-5.3135		-3.8318		-1.9196
<i>b</i> <sub>4</sub> /(MPa·K <sup>-4</sup> )			4.4400		0.5957		
<i>T</i> <sub>0</sub> /K	198.00	298.15	298.15	323.20	323.15	298.15	293.15
<i>T</i> <sub>min</sub> /K	198.00	298.15	192.79	293.15	286.80	288.15	203.00
<i>T</i> <sub>max</sub> /K	398.00	353.15	353.15	348.20	523.15	313.15	523.15
<i>P</i> <sub>min</sub> /MPa	10.00	39.23	0.12	2.00	0.57	2.50	0.35
<i>P</i> <sub>max</sub> /MPa	50.00	196.13	196.13	196.20	250.00	35.00	500.00
RMSD/(kg·m <sup>-3</sup> )	0.694	0.098	0.405	0.484	0.654	0.022	0.418
RMSD <sub>r</sub> /%	0.092	0.011	0.050	0.063	0.087	0.002	0.053
bias/(kg·m <sup>-3</sup> )	-0.018	0.002	0.052	-0.173	0.024	-0.001	0.020
<i>N</i> <sub>p</sub>	44	20	200	46	708	35	122
±	0	4	22	-12	-46	-3	16
<i>s</i> <sub>w</sub>	0.923	0.982	0.862	0.841	0.966	0.945	0.905
ref( <i>P</i> <sub>sat</sub> )			83-mcg	83-mcg	83-mcg		83-mcg

	cycloheptane	ethylcy- clohexane	1- <i>cis</i> -2-dimethyl- cyclooctane	cyclooctane	<i>trans</i> -bicyclo[4.4.0]decane	butylcyclohexane
<i>c</i> <sub>0</sub>	0.090935	0.110464	0.217447	0.086604	0.092138	0.100791
<i>b</i> <sub>0</sub> /MPa	103.3619	127.7228	154.2895	98.4214	125.1795	113.3482
<i>b</i> <sub>1</sub> /(MPa·K <sup>-1</sup> )	-62.9119	-116.7650	-72.7266	-63.1555	-56.0076	
<i>b</i> <sub>2</sub> /(MPa·K <sup>-2</sup> )	4.2958	67.1444		9.0089	-8.5697	
<i>b</i> <sub>3</sub> /(MPa·K <sup>-3</sup> )		-27.1569				
<i>b</i> <sub>4</sub> /(MPa·K <sup>-4</sup> )		4.6173				
<i>T</i> <sub>0</sub> /K	293.65	293.15	325.20	313.65	298.15	293.15
<i>T</i> <sub>min</sub> /K	293.65	293.15	325.20	313.65	298.15	293.15
<i>T</i> <sub>max</sub> /K	393.15	523.15	461.20	393.85	353.15	293.15
<i>P</i> <sub>min</sub> /MPa	5.10	0.40	0.40	5.10	39.23	2.00
<i>P</i> <sub>max</sub> /MPa	196.13	50.00	9.70	40.10	196.13	10.00
RMSD/(kg·m <sup>-3</sup> )	0.330	0.842	0.625	0.211	0.019	0.003
RMSD <sub>r</sub> /%	0.041	0.117	0.082	0.026	0.002	0.000
bias/(kg·m <sup>-3</sup> )	0.135	0.072	-0.043	-0.003	0.000	0.000
<i>N</i> <sub>p</sub>	102	93	21	75	20	5
±	40	9	-2	-3	2	1
<i>s</i> <sub>w</sub>	0.946	0.987	0.041	0.833	0.898	0.129
ref( <i>P</i> <sub>sat</sub> )		79-dyk/rep	83-mcg	84-bou/fri		

	bicyclohexyl	octadecahydro- chrysene	DPPH <sup>e</sup>	1-(1-decahydro- naphthyl)pentadecane	CPPD <sup>f</sup>	CHHU <sup>g</sup>
<i>c</i> <sub>0</sub>	0.132677	0.084675	0.090467	0.087100	0.092672	0.092218
<i>b</i> <sub>0</sub> /MPa	204.0178	117.2289	88.6294	84.8348	103.1264	88.5456
<i>b</i> <sub>1</sub> /(MPa·K <sup>-1</sup> )	-89.8586	-58.9065	-44.0926	-46.9197	-54.0928	-51.7083
<i>b</i> <sub>2</sub> /(MPa·K <sup>-2</sup> )		7.4154	17.6652	10.3755	14.6027	7.6328
<i>T</i> <sub>0</sub> /K	298.20	388.15	408.15	408.15	372.05	408.15
<i>T</i> <sub>min</sub> /K	298.20	310.93	310.93	333.15	310.95	310.95
<i>T</i> <sub>max</sub> /K	338.20	408.15	408.15	408.15	408.15	408.15
<i>P</i> <sub>min</sub> /MPa	0.69	20.00	34.46	34.46	34.46	34.46
<i>P</i> <sub>max</sub> /MPa	34.48	340.00	1033.65	585.74	1033.65	895.83
RMSD/(kg·m <sup>-3</sup> )	0.199	0.359	0.741	0.445	1.209	0.869
RMSD <sub>r</sub> /%	0.023	0.036	0.077	0.049	0.128	0.093
bias/(kg·m <sup>-3</sup> )	-0.039	-0.025	0.052	0.025	0.070	0.067
<i>N</i> <sub>p</sub>	21	48	113	49	149	103
±	-3	-6	-11	11	3	11
<i>s</i> <sub>w</sub>	0.993	0.257	0.553	0.352	0.914	0.664

	9-(2-cyclohexyl- ethyl)heptadecane	9-(3-cyclopentyl- propyl)heptadecane	DHNU <sup>h</sup>	1,3-cyclohexadiene	1,4-cyclohexadiene	cyclohexene
<i>c</i> <sub>0</sub>	0.089788	0.092100	0.081991	0.068186	0.088429	0.053518
<i>b</i> <sub>0</sub> /MPa	74.5186	94.3384	100.3959	69.6306	101.2434	53.7914
<i>b</i> <sub>1</sub> /(MPa·K <sup>-1</sup> )	-44.2218	-51.2070	-43.0062	-63.7668	-127.0709	-52.3852
<i>b</i> <sub>2</sub> /(MPa·K <sup>-2</sup> )	11.8896		21.8428			16.0062
<i>T</i> <sub>0</sub> /K	408.15	372.05	388.15	295.00	295.00	293.15
<i>T</i> <sub>min</sub> /K	310.95	310.95	310.93	293.15	293.15	293.15
<i>T</i> <sub>max</sub> /K	408.15	408.15	408.15	298.15	298.15	353.15
<i>P</i> <sub>min</sub> /MPa	34.46	34.46	20.00	1.00	1.00	1.00
<i>P</i> <sub>max</sub> /MPa	1033.65	1033.65	300.00	5.00	5.00	250.00
RMSD/(kg·m <sup>-3</sup> )	0.501	0.563	0.589	0.002	0.002	2.480
RMSD <sub>r</sub> /%	0.054	0.062	0.062	0.000	0.000	0.298
bias/(kg·m <sup>-3</sup> )	0.028	0.048	0.094	0.000	0.000	-1.117
<i>N</i> <sub>p</sub>	126	139	35	10	10	50
±	0	-17	7	0	0	-30
<i>s</i> <sub>w</sub>	0.390	0.447	0.447	0.030	0.026	1.367

<sup>a</sup> The low limit of the pressure ranges is 0.1 MPa or a saturation pressure (whichever is higher) for all fits; *P*<sub>min</sub> is the lowest pressure in a particular set of compressed liquid density data retained for the correlation. <sup>b</sup> Narrow temperature range fit. <sup>c</sup> HMTC = 2,6,10,15,19,23-hexamethyltetracosane. <sup>d</sup> HMTE = 2,6,10,15,19,23-hexamethyl-2,6,10,14,18,22-tetracosahexaene. <sup>e</sup> DPPH = 1,7-dicyclopentyl-4-(3-cyclopentylpropyl) heptane. <sup>f</sup> CPPD = 1-cyclopentyl-4-(3-cyclopentylpropyl)dodecane. <sup>g</sup> CHHU = 1-cyclohexyl-3-(2-cyclohexylethyl)undecane. <sup>h</sup> DHNU = 1,1-bis(decacyclo-1-naphthyl)undecane.



**Table 4. Statistical Characteristics of Individual Data Sets for the Fits in Table 3: Temperature and Pressure Ranges Taken into the Correlations,  $T_{\min}$ ,  $T_{\max}$ ,  $P_{\min}$ , and  $P_{\max}$ , Absolute, RMSD, and Relative,  $\text{RMSD}_r$ , Root Mean Square Deviations, Biases, bias, Number of Data Points,  $N_p$ ,  $\pm$ , and Origin of the Reference Density Values Used in the Correlations,  $\text{RD}^a$** 

ref	$T_{\min}/\text{K}$	$T_{\max}/\text{K}$	$P_{\min}/\text{MPa}$	$P_{\max}/\text{MPa}$	$\text{RMSD}/\text{kg}\cdot\text{m}^{-3}$	$\text{RMSD}_r/\%$	bias/ $\text{kg}\cdot\text{m}^{-3}$	$N_p$	$\pm$	$\text{RD}^a$
2-Methylbutane										
31-bri					4.819	0.579	3.560	9	9	o
54-isa/li	0 0									
69-mop	223.15	298.15	10.2	202.8	0.382	0.054	0.011	37	1	o
71-hou/hey					5.155	0.599	-4.962	15	-15	o
74-hou	295.15	295.15	200.0	2400.0	0.675	0.077	-0.571	12	-10	o
76-sah/gag	293.15	293.15	1.0	7.0	0.015	0.002	-0.002	7	-1	o
92-wal/bar					4.861	0.400	-4.914	88	-86	o
2,2-Dimethylpropane										
68-gon/lee					6.122	1.041	2.508	12	6	p
73-daw/sil	343.15	393.15	0.6	31.6	0.322	0.064	0.100	61	25	p
73-koh/luk	298.15	298.15	0.3	6.9	0.022	0.004	0.009	38	6	o
75-luk/dav	298.15	298.15	0.3	6.6	0.011	0.002	-0.008	49	-33	o
2-Methylpentane										
31-bri	273.15	368.15	196.2	1176.9	4.994	0.587	-0.174	26	4	o
40-kel/fel	373.15	473.15	0.6	31.6	0.912	0.198	-0.015	37	1	p
76-sah/gag	293.15	298.15	1.0	7.0	0.041	0.006	0.031	14	10	o
88-mor/aon	298.15	313.15	1.8	135.8	0.189	0.028	-0.020	64	-4	o
3-Methylpentane										
31-bri					5.892	0.827	5.785	5	5	o
52-day/fel	353.15	473.15	0.6	31.6	0.656	0.127	0.109	63	11	p
76-sah/gag	293.15	298.15	1.0	7.0	0.102	0.015	-0.089	14	-14	o
88-mor/aon	298.15	313.15	2.5	145.4	0.156	0.022	-0.001	63	-5	o
2,2-Dimethylbutane										
31-bri					71.770	10.155	-55.372	11	-11	o
43-fel/wat	373.15	473.15	1.0	30.4	0.740	0.144	-0.098	66	-16	p
76-sah/gag	293.15	298.15	1.0	7.0	0.088	0.014	0.066	14	12	o
88-mor/aon	298.15	313.15	1.5	124.3	0.271	0.039	-0.048	58	0	o
90-pol/wei	243.70	313.20	5.0	200.0	0.393	0.052	0.070	72	24	o
2,3-Dimethylbutane										
31-bri					4.188	0.579	3.704	5	5	o
42-kel/fel	373.15	473.15	0.6	31.6	0.651	0.114	0.010	37	1	o
76-sah/gag	293.15	298.15	1.0	7.0	0.159	0.024	-0.145	14	-14	o
87-hol/goe	293.15	293.15	2.0	10.0	0.041	0.006	0.017	5	1	o
88-mor/aon	298.15	313.15	1.6	36.8	0.265	0.040	-0.013	22	-4	o
90-rie/sch					4.266	0.608	-0.976	140	-40	o
93-bao/cac	208.16	298.15	0.2	106.3	0.478	0.068	-0.220	202	-88	o
2,3-Dimethylbutane (Narrow Temperature Range Fit)										
31-bri								0	0	
42-kel/fel								0	0	
76-sah/gag	293.15	298.15	1.0	7.0	0.125	0.019	-0.108	14	-14	o
87-hol/goe	293.15	293.15	2.0	10.0	0.012	0.002	0.009	5	3	o
88-mor/aon	298.15	313.15	1.6	139.3	0.208	0.030	-0.040	61	-9	o
90-rie/sch					4.831	0.688	-4.561	26	-26	o
93-bao/cac					2.481	0.349	-2.047	51	-49	o
3-Ethylpentane										
70-kus/tas	298.15	353.15	39.2	196.1	0.128	0.017	0.000	20	0	o
2,2-Dimethylpentane										
76-sah/gag	293.15	298.15	1.0	7.0	0.005	0.001	0.001	14	2	o
2,2,3-Trimethylbutane										
70-kus/tas	298.15	353.15	39.2	156.9	0.082	0.011	-0.001	11	1	o
2,2,4-Trimethylpentane										
32-bri	273.15	323.15	392.3	882.6	1.290	0.145	0.471	8	2	o
43-fel/wat-1	373.15	523.15	0.5	30.4	0.714	0.134	0.218	69	31	p
85-dym/isd	298.19	372.95	25.4	539.8	0.552	0.070	0.139	43	11	o
86-hol/goe	293.15	293.15	2.0	10.0	0.014	0.002	0.010	5	5	o
87-led					3.426	0.438	2.884	70	70	o
90-mal/woo	278.15	313.15	2.5	280.0	0.236	0.031	-0.112	92	-56	o
90-mal/woo					1.220	0.163	-1.130	46	-46	o
90-pol/wei	273.15	348.15	5.0	200.0	0.683	0.087	-0.303	48	-14	o
91-pap/zia					0.918	0.130	-0.823	9	-9	o
92-naz/gas					7.699	1.313	-3.564	0	-60	op
93-mal/woo	313.15	353.15	2.6	374.3	0.584	0.076	-0.160	89	-19	o
94-pad/far	298.15	348.15	10.1	100.1	0.648	0.090	0.544	30	30	o
96-hah/ulc	293.15	293.15	2.0	10.0	0.018	0.003	-0.010	5	-3	o
96-pad/far-1	197.93	248.35	10.1	100.1	1.039	0.134	-0.663	27	-17	o
96-pad/far-1	298.15	348.16	10.1	100.1	0.833	0.114	-0.683	30	-26	o
Heptadecane										
64-doo	323.15	573.15	5.0	500.0	1.136	0.145	0.054	60	-2	o
87-man/cri	298.15	338.15	2.0	10.1	0.071	0.009	-0.006	27	-9	o

**Table 4. Continued**

ref	$T_{\min}/K$	$T_{\max}/K$	$P_{\min}/MPa$	$P_{\max}/MPa$	RMSD/kg·m <sup>-3</sup>	RMSD <sub>r</sub> /%	bias/kg·m <sup>-3</sup>	$N_p$	±	RD <sup>a</sup>
					Octadecane					
58-cut/mcm	333.15	408.15	34.5	551.3	0.403	0.051	0.023	48	-4	o
					7-Hexyltridecane					
59-low/spe	310.93	408.15	20.0	340.0	0.230	0.028	0.027	36	-2	o
					Eicosane					
64-doo	373.15	573.15	5.0	500.0	1.222	0.159	0.067	50	0	o
					Docosane					
88-pet/spi	323.08	368.26	2.1	16.1	0.422	0.055	-0.001	48	-16	p
					Tetracosane					
87-pet/van	333.52	371.22	2.1	12.1	0.195	0.026	0.000	36	-12	p
					9-Octylheptadecane					
58-cut/mcm	310.95	408.15	34.5	1033.7	0.640	0.072	-0.051	116	-16	o
59-low/spe	310.93	372.04	20.0	340.0	0.737	0.089	0.703	20	20	o
					Triacontane					
64-doo	373.15	573.15	5.0	500.0	0.973	0.124	0.099	47	11	o
					2,6,10,15,19,23-Hexamethyltetracosane					
70-kus/tas	298.15	353.15	39.2	196.1	0.057	0.007	0.000	20	0	o
					11-Decylheneicosane					
59-low/spe	310.93	408.15	40.0	340.0	0.302	0.036	0.022	31	1	o
					13-Dodecylhexacosane					
59-low/spe	310.93	408.15	20.0	340.0	0.226	0.027	0.029	30	4	o
					Tetracontane					
64-doo	423.15	573.15	5.0	500.0	1.267	0.159	0.142	40	8	o
					2-Methyl-1,3-butadiene					
32-bri	273.15	273.15	49.0	1176.8	0.568	0.064	0.041	14	2	o
					1-Pentene					
51-day/fel	353.15	448.15	0.6	31.6	0.453	0.090	-0.056	50	-20	op
					2-Methyl-2-butene					
49-bri	298.15	298.15	49.0	490.3	0.205	0.027	0.014	10	-4	o
					(E)-1,4-Hexadiene					
75-bur/ric	293.15	298.15	1.0	5.0	0.002	0.000	0.000	10	2	o
					(2Z,4E)-2,4-Hexadiene					
75-bur/ric	293.15	298.15	1.0	5.0	0.002	0.000	0.000	10	2	o
					1,5-Hexadiene					
75-bur/ric	293.15	298.15	1.0	5.0	0.001	0.000	0.000	10	-2	o
					1-Hexene					
70-aba/ker					1.279	0.199	-0.492	160	-62	op
72-ker/apa	283.15	473.15	0.8	68.7	1.060	0.167	-0.262	366	-102	op
75-bur/ric	293.15	298.15	1.0	5.0	0.024	0.004	0.015	10	4	o
82-gus/gal	146.00	274.00	5.0	50.0	0.672	0.088	-0.168	42	-8	o
					1-Heptene					
81-gus/naz	293.15	353.15	5.0	50.0	1.162	0.169	-0.165	23	-7	o
					1-Octene					
49-bri					3.440	0.430	-3.389	6	-6	o
72-ker/apa					3.186	0.486	0.709	361	78	ope
88-dym/mal	298.15	373.15	10.6	312.7	0.616	0.082	-0.076	33	-1	o
92-naz/gas	290.15	538.95	5.0	58.9	1.576	0.244	0.516	67	15	op
					1-Nonene					
84-gus/gal	198.00	398.00	10.0	50.0	0.694	0.092	-0.018	44	0	o
					2,6,10,15,19,23-Hexamethyl-2,6,10,14,18,22-tetracosahexaene					
70-kus/tas	298.15	353.15	39.2	196.1	0.098	0.011	0.002	20	4	o
					Cyclopentane					
69-bra/fre					3.698	0.452	3.402	3	3	o
70-kus/tas	313.15	353.15	39.2	196.1	1.503	0.185	0.682	14	6	op
92-bao/cac	192.79	298.15	0.1	104.3	0.081	0.010	0.005	186	16	o
					Methylcyclopentane					
80-oza/ooy	298.20	348.20	9.9	196.2	0.512	0.066	-0.194	41	-11	op
87-hol/goe	293.15	293.15	2.0	10.0	0.010	0.001	-0.002	5	-1	o
					Cyclohexane					
34-rot/nag					0.538	0.068	0.467	9	9	o
40-sch/hof					1.753	0.241	0.308	3	1	e
57-rea/sag					2.396	0.357	1.077	78	40	o
59-gol/vag					3.959	0.697	1.793	64	16	e
59-gol/vag					7.709	1.226	-5.683	56	-56	op
62-hol/wha	298.16	348.15	1.0	10.0	0.008	0.001	0.005	50	34	o

Table 4. Continued

ref	$T_{\min}/\text{K}$	$T_{\max}/\text{K}$	$P_{\min}/\text{MPa}$	$P_{\max}/\text{MPa}$	RMSD/kg·m <sup>-3</sup>	RMSD <sub>r</sub> /%	bias/kg·m <sup>-3</sup>	$N_p$	$\pm$	RD <sup>a</sup>
Cyclohexane (continued)										
70-kus/tas					1.565	0.201	-1.492	6	-6	o
72-gol/ada					6.137	0.763	-5.887	24	-24	o
72-ker/apa					1.636	0.242	-1.160	440	-356	op
72-lys					15.999	1.875	-15.999	1	-1	o
73-ker/apa					1.455	0.215	-1.107	272	-212	op
73-rog/bur	298.15	298.15	1.0	10.0	0.075	0.010	-0.064	10	-10	o
74-apa/ker					2.707	0.426	-1.572	90	-56	op
75-bur/ric	293.15	298.15	1.0	5.0	0.029	0.004	-0.025	10	-10	o
75-gri/mur	498.15	523.15	2.1	79.1	1.046	0.185	0.322	28	14	p
75-ras/gri	298.15	473.15	0.6	85.4	0.271	0.038	0.013	96	-6	op
75-ras/gri	298.15	473.15	1.6	71.0	0.121	0.017	-0.001	35	1	op
78-gou	314.15	393.15	5.1	40.1	0.428	0.057	0.164	75	37	p
79-dic								0	0	e
79-isd/dym	298.15	348.15	10.0	20.0	0.272	0.035	-0.245	4	-4	o
79-isd/dym	348.15	348.15	50.0	100.0	0.869	0.110	-0.868	2	-2	o
79-kas/fuk	298.15	348.15	6.9	105.3	0.313	0.040	-0.192	32	-18	o
80-jon/has	313.00	383.00	5.0	214.0	0.947	0.122	-0.659	38	-26	ope
82-wis/wue	286.80	337.90	10.0	110.0	0.496	0.063	-0.219	24	-6	o
84-mat/van	288.15	313.15	2.5	35.0	0.120	0.015	0.094	34	34	o
87-sun/kor	288.15	323.06	5.0	85.0	0.260	0.033	-0.233	80	-80	o
88-mor/aon	298.15	313.15	2.4	46.1	0.216	0.028	-0.098	27	-9	o
89-vos/slo					3.755	0.563	3.299	24	24	e
90-pol/wei	293.20	368.10	5.0	180.0	0.793	0.097	0.538	35	21	o
90-tos/fig					1.787	0.227	-1.483	48	-40	o
90-tos/fig					1.680	0.211	-1.579	6	-6	o
91-mel/mel	293.15	413.15	10.0	250.0	1.855	0.234	0.850	49	13	pe
91-tan/hos	298.15	348.15	6.2	100.0	0.264	0.034	-0.202	23	-21	o
92-lag/bon	288.15	323.15	5.0	15.0	0.334	0.043	-0.249	18	-14	o
92-lag/bon	288.15	323.15	5.0	15.0	0.331	0.042	0.229	18	14	o
96-pad/far	298.15	348.14	5.1	38.0	0.359	0.047	-0.140	20	-8	o
Cyclohexane- $d_{12}$										
84-mat/van	288.15	313.15	2.5	35.0	0.022	0.002	-0.001	35	-3	o
Methylcyclohexane										
49-bri					4.330	0.519	-3.459	10	-8	o
69-bra/fre					1.899	0.216	0.287	9	-1	o
70-aba/ker					3.347	0.490	-2.415	195	-185	op
72-ker/apa					4.305	0.628	-2.446	351	-331	op
72-ker/apa	283.15	523.15	0.3	4.0	0.433	0.068	-0.024	20	-4	op
78-gou	293.15	392.55	5.1	40.1	0.191	0.025	0.075	75	25	p
79-jon/has	203.00	298.00	50.0	500.0	0.819	0.094	-0.083	21	-3	o
97-bay/bon	303.15	343.15	20.0	40.0	0.351	0.046	-0.157	6	-2	o
Cycloheptane										
70-kus/tas	298.15	353.15	39.2	196.1	0.211	0.026	-0.100	12	-4	o
78-gou	293.65	393.15	5.1	40.1	0.343	0.043	0.166	90	44	p
Ethylcyclohexane										
83-gus/sha	293.15	523.15	5.0	50.0	0.781	0.111	-0.054	72	-4	op
89-vos/slo	327.10	440.90	0.4	9.7	1.022	0.136	0.502	21	13	p
1- <i>cis</i> -2-Dimethylcyclohexane										
89-vos/slo	325.20	461.20	0.4	9.7	0.625	0.082	-0.043	21	-2	p
Cyclooctane										
78-gou	313.65	393.85	5.1	40.1	0.211	0.026	-0.003	75	-3	p
<i>trans</i> -Bicyclo[4.4.0]decane ( <i>trans</i> -decalin)										
70-kus/tas	298.15	353.15	39.2	196.1	0.019	0.002	0.000	20	2	o
Butylcyclohexane										
87-hol/goe	293.15	293.15	2.0	10.0	0.003	0.000	0.000	5	1	o
Bicyclohexyl										
88-sid/tej	298.20	338.20	0.7	34.5	0.199	0.023	-0.039	21	-3	o
Octadecahydrochrysene										
59-low/spe	310.93	408.15	20.0	340.0	0.359	0.036	-0.025	48	-6	o
1,7-Dicyclopentyl-4-(3-cyclopentylpropyl)heptane										
58-cut/mcm	310.95	408.15	34.5	1033.7	0.741	0.077	0.052	113	-11	o
1-(1-Decahydronaphthyl)pentadecane										
58-cut/mcm	333.15	408.15	34.5	585.7	0.445	0.049	0.025	49	11	o
1-Cyclopentyl-4-(3-cyclopentylpropyl)dodecane										
58-cut/mcm	310.95	408.15	34.5	1033.7	1.209	0.128	0.070	149	3	o
1-Cyclohexyl-3-(2-cyclohexylethyl)undecane										
58-cut/mcm	310.95	408.15	34.5	895.8	0.869	0.093	0.067	103	11	o

Table 4. Continued

ref	$T_{\min}/\text{K}$	$T_{\max}/\text{K}$	$P_{\min}/\text{MPa}$	$P_{\max}/\text{MPa}$	RMSD/kg·m <sup>-3</sup>	RMSD <sub>r</sub> /%	bias/kg·m <sup>-3</sup>	$N_p$	±	RD <sup>a</sup>
58-cut/mcm	310.95	408.15	34.5	1033.7	0.501	0.054	0.028	126	0	o
58-cut/mcm	310.95	408.15	34.5	1033.7	0.563	0.062	0.048	139	-17	o
59-low/spe	310.93	408.15	20.0	300.0	0.589	0.062	0.094	35	7	o
75-bur/ric 79-dic	293.15	298.15	1.0	5.0	0.002	0.000	0.000	10 0	0 0	o
75-bur/ric 79-dic	293.15	298.15	1.0	5.0	0.002	0.000	0.000	10 0	0 0	o
72-ker/apa 75-bur/ric 79-dic	293.15	298.15	1.0	5.0	0.012	0.001	0.001	10 0	-2 0	o
91-mel/mel	293.15	353.15	10.0	250.0	2.773	0.333	-1.396	40	-28	p

<sup>a</sup> o, from the same source as the compressed liquid density data; e, from the smoothing equation (see Appendix); p, extrapolated to  $P_{\text{ref}}$  along isotherms using the Tait equation.

The temperature and pressure ranges of the validity of the fits given in the table allow one to avoid extrapolation using eq 1 with the parameters from Table 3 beyond the  $P$ - $T$  areas of the retained data. The  $P$ - $T$  areas that are not rectangular are shown in the figures (Figure 1), which provide crude information on the distribution of the retained data points. Nonrectangular  $P$ - $T$  areas appeared mostly for measurements where the  $P$ - $T$  range approached the vicinity of a solid-liquid equilibrium line.

Table 4 summarizes some statistical information derived from the fits. Only those data subsets for which the temperature and pressure ranges are displayed in the table were retained in the correlations. The statistical characteristics of these subsets refer only to the data points retained in the correlation. On the other hand, the characteristics of the rejected subsets, that is, those for which no  $T$  and  $P$  ranges are given in the table, illustrate the deviations of the rejected points from eq 1, but only for those values within the  $P$ - $T$  areas of the retained data (see Table 3 and Figure 1). If only one set of smoothed  $P$ - $\rho$ - $T$  values available in original sources (S-type data) or one set of values generated from published equations (F-type data) for a particular substance was fitted by eq 1, then the average deviations of the fit (RMSD, RMSD<sub>r</sub>) do not reflect the real accuracy of the experiment and consequently the weighted standard deviation  $s_w$  may be lower than unity (the deviations of both the smoothed input data and the F-type data from eq 1 are lower than the experimental uncertainty declared by the authors).

Values at high temperatures of some retained data sets were rejected in those cases where large deviations from the Tait equation were observed and it was not possible to improve the fit by the additional parameters  $b_i$  and  $c_i$ . Thus, the  $P$ - $T$  ranges of some fits do not cover the entire original range of retained data sets. The temperature and/or pressure ranges were sometimes enlarged by retaining less accurate and less reliable values in the ranges beyond those of more accurate data sets but only in those cases where the representation of accurate data was not affected by the enlargement and the enlargement did not result in a distortion of the  $B(T)$  function. The absence of extremes and inflection points on the function  $B(T)$  (eq 3) of all final fits was checked.

A comparison of isothermal compressibilities,  $\beta_T = -(1/V)(\partial V/\partial P)_T = (1/\rho)(\partial \rho/\partial P)_T$ , calculated from the fits for  $P =$

0.1 MPa with available values published in the literature is presented in Table 5, which provides a rough check of the consistency of the fits with independent data. The literature values of isothermal compressibility used for the comparison in Table 5 are the values obtained mostly from speed-of-sound measurements and were either taken directly from the papers or calculated from the equation

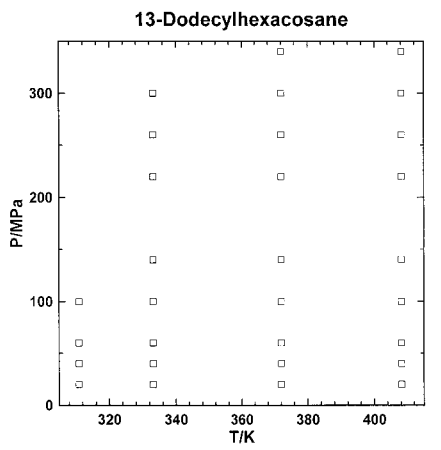
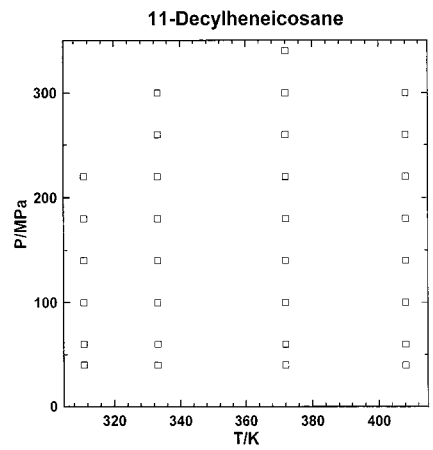
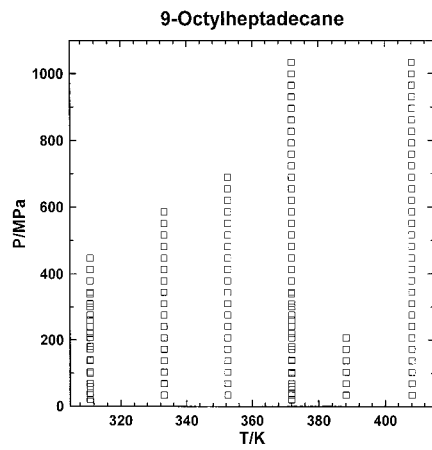
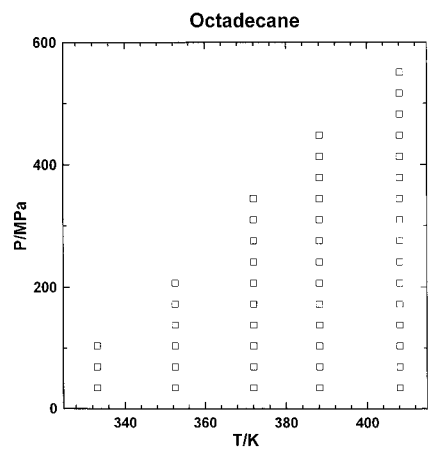
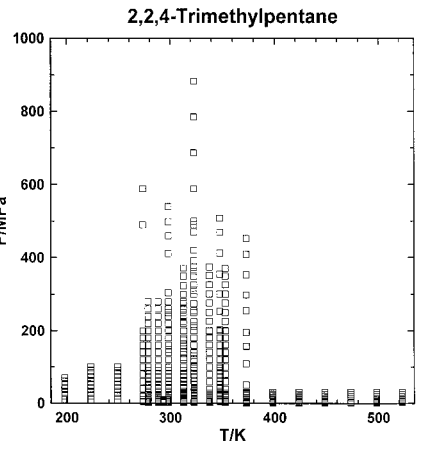
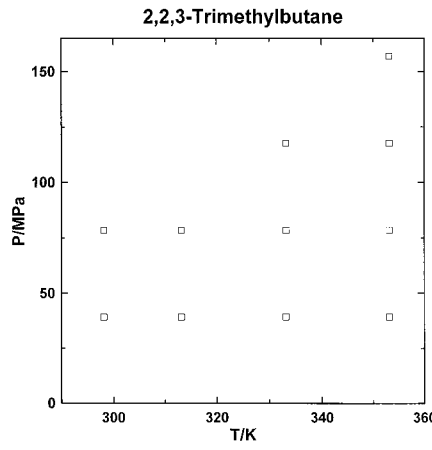
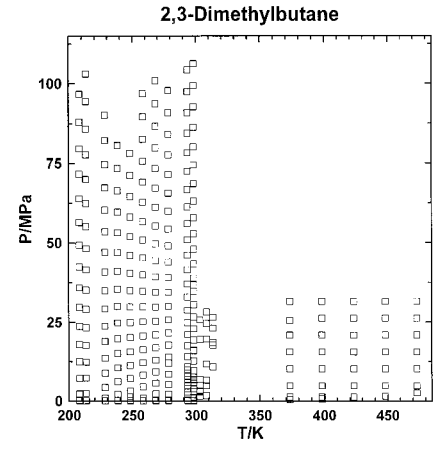
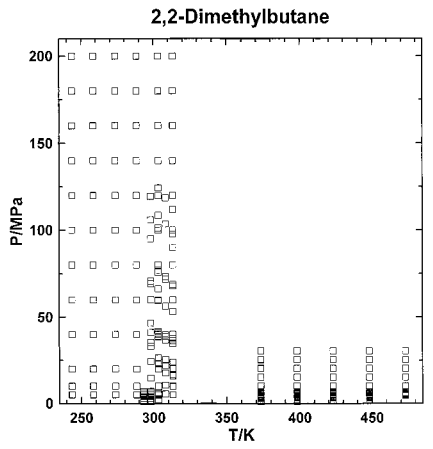
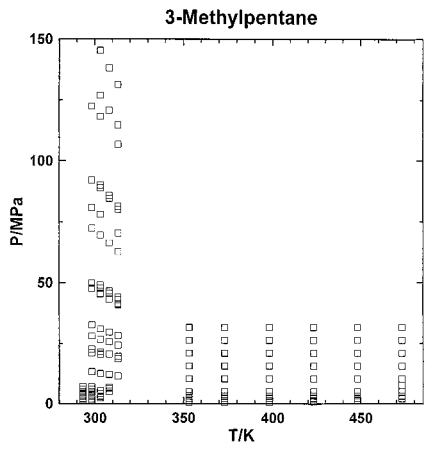
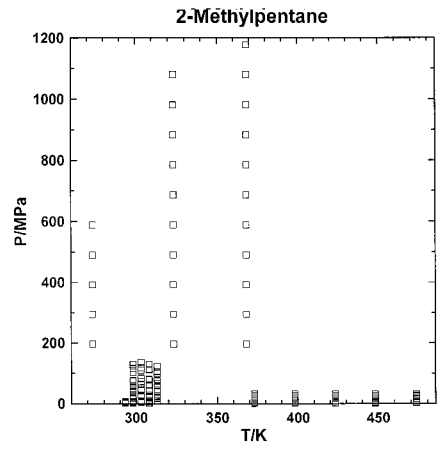
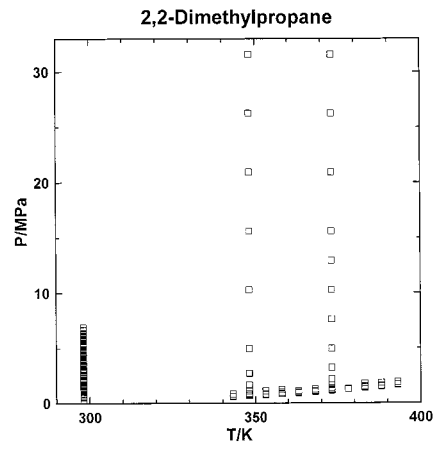
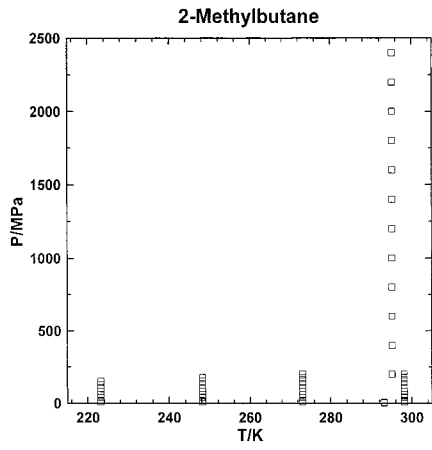
$$\beta_T = \frac{1}{\rho} \left[ \frac{1}{u^2} + \frac{TM\alpha_p^2}{c_p} \right] \quad (11)$$

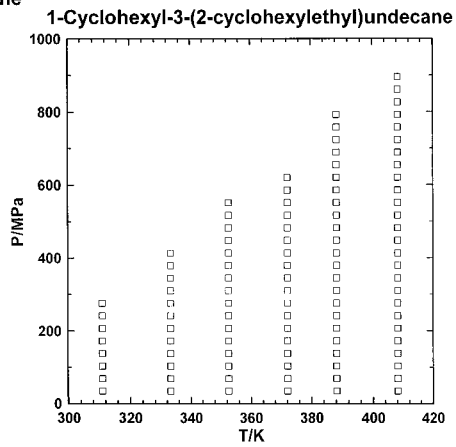
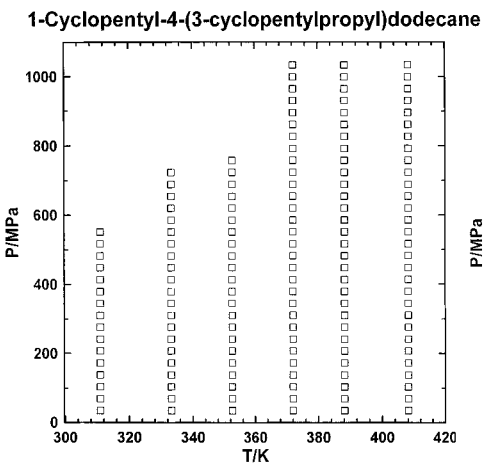
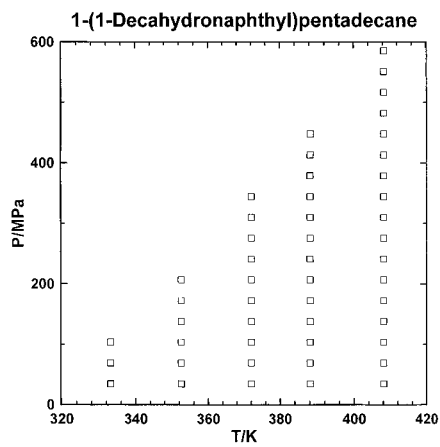
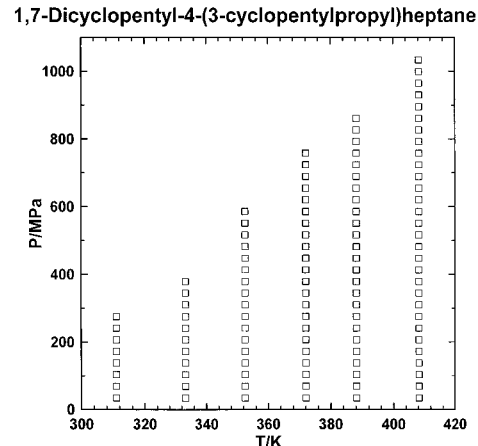
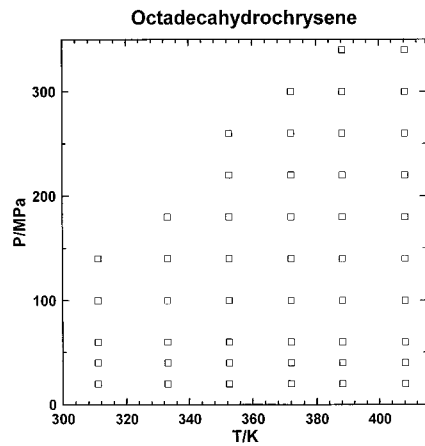
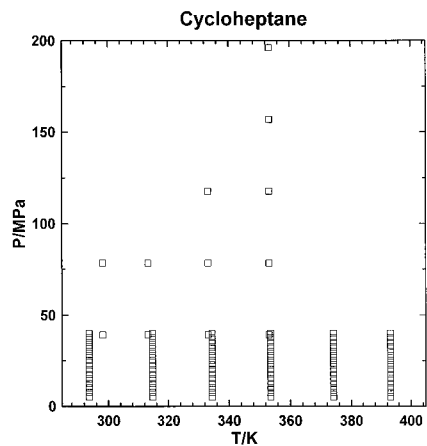
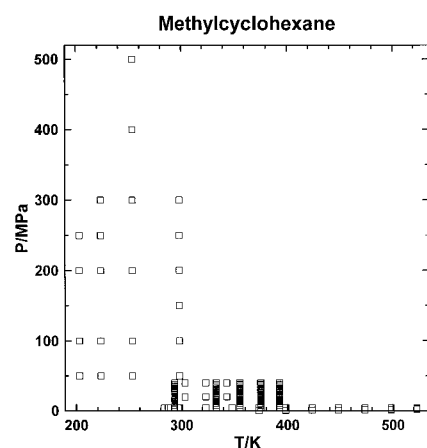
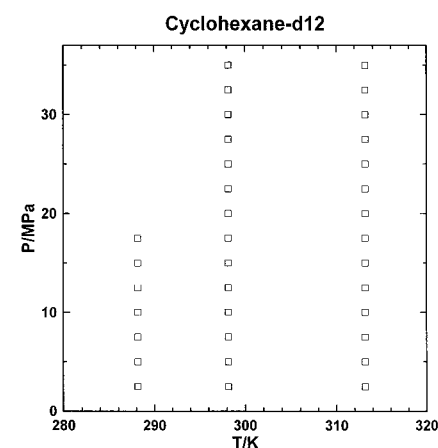
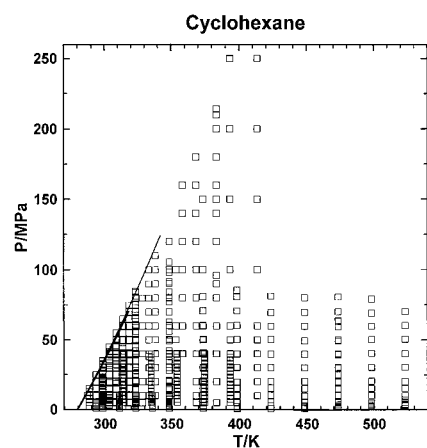
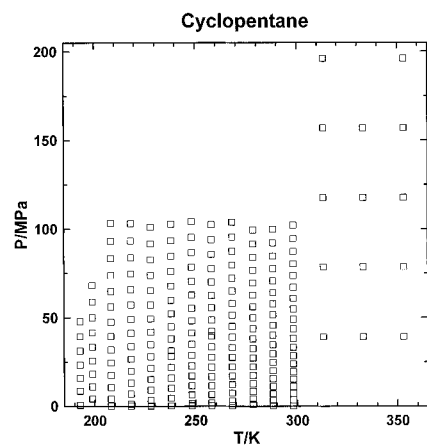
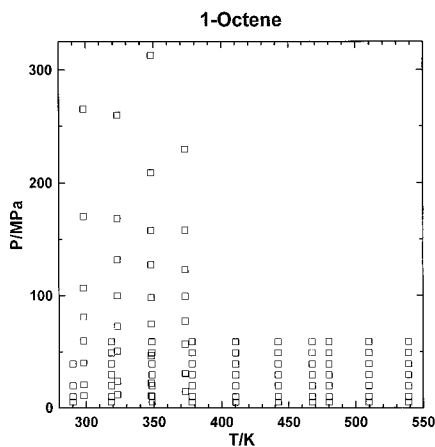
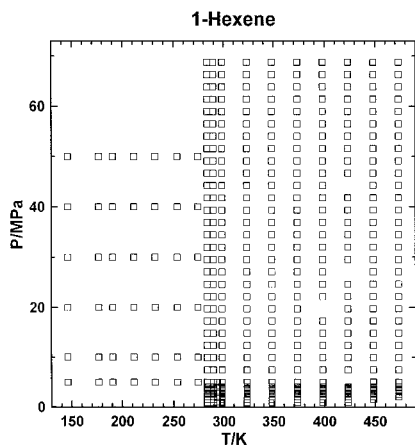
where  $M$ ,  $u$ ,  $\alpha_p$ , and  $c_p$  are the molar mass, speed of sound, isobaric thermal expansivity ( $\alpha_p = (1/V)(\partial V/\partial T)_p = -(1/\rho)(\partial \rho/\partial T)_p$ ), and molar isobaric heat capacity, respectively. The values of the input quantities in eq 11 were taken from the different sources cited in Table 5.

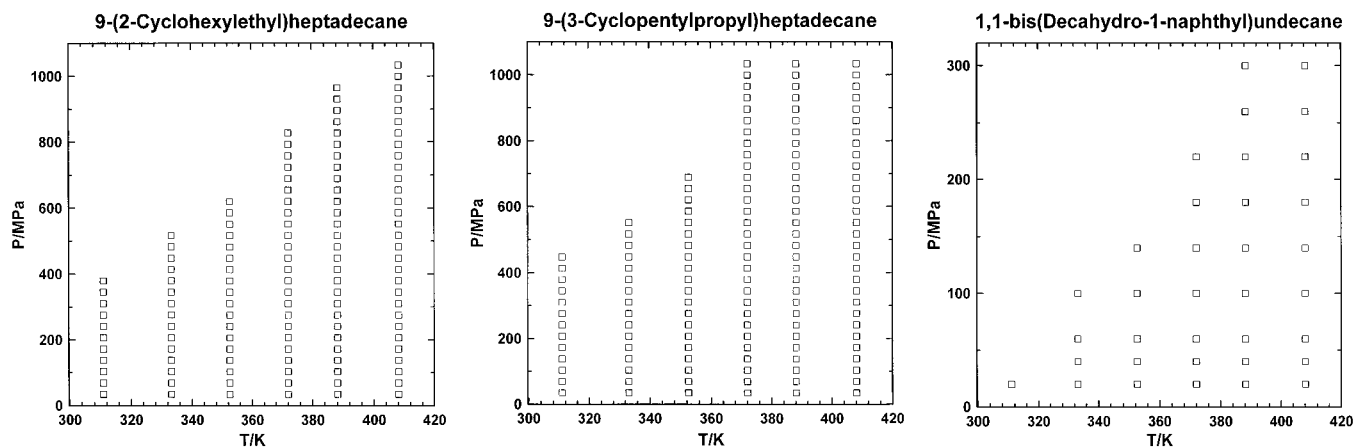
**Alkanes.** Data [92-wal/bar] for 2-methylbutane in the low-temperature range differ from those by Mopsik [69-mop]. Tentative fits of each of those data sets revealed that the data of Mopsik [69-mop] gave much better agreement (positive deviations about 1% on average) with isothermal compressibilities evaluated from speed of sound [92-wal/bar] than with  $P$ - $\rho$ - $T$  data [92-wal/bar] (negative deviations in the range from -1.6% at 203 K to -12.2% at 243 K). The final fit yields the average deviation in the isothermal compressibility 2% (positive, see Table 5), even when the values extrapolated beyond the temperature range of the fit are included.

The agreement between the calculated and experimental isothermal compressibilities of 2-methylpentane is satisfactory (deviation 1.8% on average) at temperatures where the fit represents the data of Sahli et al. and Moriyoshi and Aono [76-sah/gag, 88-mor/aon] while a larger deviation (-4.3%) is observed at 273.15 K where the fit is based on less accurate data by Bridgman [31-bri]. Bridgman's data were, however, rejected for  $P < 196$  MPa.

The average deviations of the data retained for 3-methylpentane and 2,2-dimethylbutane from the final fits (Table 3) are within 0.1%, and the average deviations of the isothermal compressibilities (Table 5) are 0.5 and 1.9%, respectively. The deviation of the isothermal compressibility extrapolated from the fit for 3-methylpentane is 5.1% at 273.15 K (20 K below  $T_{\min}$  of the fit); much better agreement (1.4%) is observed for  $T = 283.15$  K.







**Figure 1.** Temperature and pressure coordinates of data points retained in the correlations for the fits in Table 3 where  $P$ - $T$  areas of retained data points are not rectangular. The lines represent solid-liquid equilibrium curves; cyclohexane (normal melting temperature 279.6 K): [75-ras/gri] {up to 318 K (70 MPa)}, [82-wis/wue] {up to 338 K (117 MPa)}, [93-yok/ebi] {up to 341 K (122 MPa)}.

Two fits are presented for 2,3-dimethylbutane. The first one, performed in as wide a temperature range as possible, yielded isothermal compressibilities 2.4% higher than the available literature values (Table 5); the second fit, representing data [76-sah/gag, 87-hol/goe, 88-mor/aon] in the range 293.15–313.15 K, is in perfect agreement with experimental data on isothermal compressibilities. Data [90-rie/sch] exhibit large deviations and were rejected in both cases. From the comparison of both fits (Table 4), it, however, follows that the values of Baonza et al. [93-bao/cac] are in the range 293.15–298.15 K, lower than other data retained for the narrow-temperature-range fit.

Fits of the data of Kuss and Taslimi [70-kus] for 3-ethylpentane and 2,2,3-trimethylbutane yielded similar deviation from available isothermal compressibilities (within 6%). The fit for 2,2-dimethylpentane is the fit of the F-type data of Sahli et al. [76-sah/gag]; the agreement for the isothermal compressibility is very good.

The fit for 2,2,4-trimethylpentane covers a temperature range more than 300 K wide. The experimental data are in good agreement, and few data sets were rejected (Table 4). The average deviation of the isothermal compressibility in the range 233.15–318.15 K is excellent, being 1.2%. Values reported in [96-pad/far-1] for the temperature range 298–348 K represent the same “raw” data as those from [94-pad/far], but they were analyzed by Padua and co-workers using a theoretically improved model of the vibrating-wire densimeter. It was, however, found that the fit of the values of [94-pad/far] gives much better agreement with available isothermal compressibilities at 0.1 MPa than that of the values of [96-pad/far-1] (average deviation –2.5% in the range 233–318 K). Therefore, both sets of density values were retained in the correlations, each with its declared uncertainty (0.1% for [94-pad/far], 0.05% for [96-pad/far-1]).

Fits for other alkanes are mostly fits of a set of data from one source. Except for dodecane, where large deviations from literature values are observed, the data to obtain independent isothermal compressibilities for comparison were not available. The data along the isotherm 388.15 K for 9-octylheptadecane [58-cut/mcm] were rejected to obtain satisfactory consistency with the data from the same laboratory [59-low/spe]. The data along the isobars 20 MPa of 11-decylheneicosane (negative deviations 3–5  $\text{kg}\cdot\text{m}^{-3}$ ) and 180 MPa of 13-dodecylhexacosane (positive deviations 1.8–2.4  $\text{kg}\cdot\text{m}^{-3}$ ) were rejected.

**Alkenes.** The upper temperature limit of the data for 1-pentene is 448.15 K; the authors [51-day/fel] declared

that observable polymerization of the substance occurred at temperatures higher than 473 K. The data for 1-hexene cover the temperature range up to 503 K. No polymerization was mentioned by the authors, but the fit was limited up to 473.15 K, since large deviations were observed for higher temperatures. Agreement of the isothermal compressibilities calculated from the fit for 293.15 and 298.15 K is very good (Table 5).

Similarly the upper temperature limit of the fit for 1-heptene is 353 K while the data covers temperatures up to 523 K. When the temperature range of the correlation was increased, points of inflection and even local extremes appeared on the function  $B(T)$ . The isothermal compressibilities calculated from the final fit are 19% lower than experimental values. Similar results were reached in the case of the data for 1-nonene from the same laboratory [84-gus/gal]; the temperature range of the final fit was limited up to 398 K; the deviations in isothermal compressibilities are –26%. It is worth mentioning that the values of the Tait parameters  $c_b$  and  $b_f$  for these two alkenes (Table 3) seem to be rather unusual.

The fit for 1-octene represents the data of Dymond et al. [88-dym/mal] and Naziev et al. [92-naz/gas]. Rather large deviations were observed for the data of Naziev et al. [92-naz/gas], but the increase of the number of adjustable parameters led to distortion of the  $B(T)$  function. The agreement of the isothermal compressibilities calculated from the final fit with literature values is satisfactory.

**Cycloalkanes.** The fit for cyclopentane represents predominantly of data of Baonza et al. [92-bao/cac]; the temperature range of the fit was extended by retaining the data of Kuss and Taslimi [70-kus/tas]. The calculated isothermal compressibilities agree with the literature values within 2% (neglecting the less accurate value of Shinoda and Hildebrand [61-shi/hil]).

The fit for cyclohexane covers the temperature range from a few Kelvins above the normal melting point (279.6 K) up to 30 K below the critical temperature. The deviations of the calculated isothermal compressibility from literature data are very satisfactory, being about 1.5% in the temperature interval 293 to 352 K (see Table 5). Several data sets for cyclohexane present values at higher temperatures. Among them, the data of Grigoriev et al. [75-gri/mur, 75-ras/gri] were retained, since much better agreement with other retained data was observed at lower temperatures for these data than for the values of Kerimov et al. [72-ker/apa, 73-ker/apa, 74-apa/ker]. This is in accordance with the results of Penoncello et al. [95-pen/

**Table 5. Comparison of Values of Isothermal Compressibility,  $\beta_T = (1/\rho)(\partial\rho/\partial P)_T$ , at  $P = 0.1$  MPa Calculated from the Fits in Table 3 (Eq 1) with Literature Data**

$T/K$	$\beta_T/\text{GPa}^{-1}$		$\delta\beta_T/\%$ <sup>b</sup>	ref(s)	$T/K$	$\beta_T/\text{GPa}^{-1}$		$\delta\beta_T/\%$ <sup>b</sup>	ref(s)
	eq 1 <sup>a</sup>	lit.				eq 1 <sup>a</sup>	lit.		
2-Methylbutane									
203.00	0.911 <sup>f</sup>	0.893	2.0	92-wal/bar, <sup>d</sup> 96-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	233.00	1.198 ± 0.006	1.168	2.6	92-wal/bar, <sup>d</sup> 96-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
213.00	0.996 <sup>f</sup>	0.977	1.9	92-wal/bar, <sup>d</sup> 96-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	243.00	1.320 ± 0.006	1.291	2.2	92-wal/bar, <sup>d</sup> 96-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
223.00	1.092 ± 0.007	1.073	1.8	92-wal/bar, <sup>d</sup> 96-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>					
2-Methylpentane									
273.15	1.415 ± 0.012	1.479	-4.3	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.793 ± 0.007	1.807	-0.8	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
283.15	1.551 ± 0.009	1.595	-2.7	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.833	-2.2	89-ohn/fuj <sup>c</sup>
298.15	1.793 ± 0.007	1.839	-2.5	83-aic/kum, <sup>c</sup> 88-tre/ben <sup>c</sup>	313.15	2.091 ± 0.008	2.072	0.9	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
3-Methylpentane									
273.15	1.392 <sup>f</sup>	1.325	5.1	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.714 ± 0.006	1.719	-0.3	83-aic/kum, <sup>c</sup> 88-tre/ben <sup>c</sup>
283.15	1.508 <sup>f</sup>	1.486	1.4	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.699	0.9	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
293.15	1.641 ± 0.007	1.627	0.9	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.714	0.0	89-ohn/fuj <sup>c</sup>
					313.15	1.966 ± 0.007	1.970	-0.2	83-aww/pet, <sup>d</sup> 92-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
2,2-Dimethylbutane									
298.15	1.967 ± 0.008	2.014	-2.3	82-tre/han <sup>c</sup>	298.15	1.967 ± 0.008	1.997	-1.5	85-cos/bha <sup>c</sup>
		2.010	-2.1	83-aic/kum, <sup>c</sup> 88-tre/ben <sup>c</sup>			2.002	-1.7	89-ohn/fuj <sup>c</sup>
2,3-Dimethylbutane									
298.15	1.831 ± 0.006	1.790	2.3	83-aic/kum, <sup>c</sup> 88-tre/ben <sup>c</sup>	298.15	1.831 ± 0.006	1.786	2.5	89-ohn/fuj <sup>c</sup>
298.15	1.787 ± 0.006 <sup>g</sup>	1.790	-0.2	83-aic/kum, <sup>c</sup> 88-tre/ben <sup>c</sup>	298.15	1.787 ± 0.006 <sup>g</sup>	1.786	0.05	89-ohn/fuj <sup>c</sup>
3-Ethylpentane									
318.15	1.606 ± 0.007	1.701	-5.6	29-fre/hub, <sup>d</sup> 91-trc-1, <sup>e</sup> 96-zab/ruz <sup>f</sup>					
2,2-Dimethylpentane									
293.15	1.569 ± 0.004	1.604	-2.2	29-fre/hub, <sup>d</sup> 91-trc-1, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.667 ± 0.004	1.660	0.4	84-aww/pet, <sup>d</sup> 91-trc-1, <sup>e</sup> 96-zab/ruz <sup>f</sup>
2,2,3-Trimethylbutane									
298.15	1.482 ± 0.006	1.566	-5.4	88-tre/ben <sup>c</sup>	318.15	1.781 ± 0.007	1.717	3.7	29-fre/hub, <sup>d</sup> 91-trc-1, <sup>e</sup> 96-zab/ruz <sup>f</sup>
2,2,4-Trimethylpentane									
233.15	0.942 ± 0.006	0.924	1.9	83-aww/pet, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.535 ± 0.002	1.546	-0.7	91-pap/zia, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
253.15	1.089 ± 0.004	1.068	2.0	83-aww/pet, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.541	-0.4	94-ami/ara, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
263.15	1.172 ± 0.003	1.157	1.3	83-aww/pet, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.563	-1.8	95-mie/osw <sup>c</sup>
273.15	1.264 ± 0.002	1.259	0.4	83-aww/pet, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.550	-1.0	95-ami/gop, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
293.15	1.475 ± 0.002	1.471	0.3	83-aww/pet, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	308.15	1.664 ± 0.003	1.678	-0.8	94-ami/ara, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
298.15	1.535 ± 0.002	1.56	-1.6	61-shi/hil <sup>c</sup>	313.15	1.734 ± 0.003	1.774	-2.3	95-mie/osw <sup>c</sup>
		1.542	-0.5	82-tre/han, <sup>c</sup> 88-tre/ben, <sup>c</sup> 85-cos/bha <sup>c</sup>	318.15	1.808 ± 0.003	1.844	-2.0	29-fre/hub, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.543	-0.5	83-aww/pet, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.834	-1.4	94-ami/ara, <sup>d</sup> 90-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
Docosane									
323.15	1.02 ± 0.07	0.900	13.3	82-bro/lop <sup>h</sup>	353.15	1.36 ± 0.10	1.066	27.2	82-bro/lop <sup>h</sup>
333.15	1.14 ± 0.07	0.950	19.7	82-bro/lop <sup>h</sup>	363.15	1.45 ± 0.11	1.130	27.6	82-bro/lop <sup>h</sup>
343.15	1.26 ± 0.08	1.006	25.2	82-bro/lop <sup>h</sup>					
1-Hexene									
293.15	1.589 ± 0.005	1.584	0.3	58-par/pan, <sup>d</sup> 95-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.666 ± 0.005	1.689	-1.4	86-tar/dia <sup>c</sup>
298.15	1.666 ± 0.005	1.614	3.2	75-bur/ric <sup>c</sup>					
1-Heptene									
283.15	1.06 <sup>f</sup>	1.289	-17.8	48-lag/mcm, <sup>d</sup> 95-trc-1, <sup>e</sup> 96-zab/ruz <sup>f</sup>	303.15	1.27 ± 0.07	1.577	-19.5	48-lag/mcm, <sup>d</sup> 95-trc-1, <sup>e</sup> 96-zab/ruz <sup>f</sup>
293.15	1.16 ± 0.06	1.437	-19.3	48-lag/mcm, <sup>d</sup> 95-trc-1, <sup>e</sup> 96-zab/ruz <sup>f</sup>					
1-Octene									
283.15	1.215 <sup>f</sup>	1.162	4.6	48-lag/mcm, <sup>d</sup> 86-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	293.15	1.285 <sup>f</sup>	1.243	3.4	58-par/pan <sup>c</sup>
293.15	1.285 <sup>f</sup>	1.259	2.1	48-lag/mcm, <sup>d</sup> 86-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	303.15	1.361 ± 0.010	1.354	0.5	48-lag/mcm, <sup>d</sup> 86-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
1-Nonene									
283.15	0.79 ± 0.02	1.067	-26.0	48-lag/mcm, <sup>d</sup> 86-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	0.88 ± 0.02	1.182	-25.5	48-lag/mcm, <sup>d</sup> 86-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
293.15	0.85 ± 0.02	1.145	-25.8	48-lag/mcm, <sup>d</sup> 86-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	303.15	0.91 ± 0.03	1.229	-26.0	48-lag/mcm, <sup>d</sup> 86-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
Cyclopentane									
298.15	1.308 ± 0.006	1.35	-3.1	61-shi/hil <sup>c</sup>	298.15	1.308 ± 0.006	1.332	-1.8	86-tar/dia <sup>c</sup>
		1.331	-1.7	74-ewi/mar, <sup>c</sup> 77-ewi/mar <sup>c</sup>	303.15	1.369 ± 0.006	1.379	-0.7	49-wei, <sup>d</sup> 95-trc-3, <sup>e</sup> 96-zab/ruz <sup>f</sup>
Cyclohexane									
292.85	1.086 ± 0.002	1.051	3.3	85-lav/jak, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.135 ± 0.002	1.113	2.0	95-ami/gop, <sup>d,e</sup> 96-zab/ruz <sup>f</sup>
293.15	1.089 ± 0.002	1.038	4.9	52-jac <sup>c</sup>	303.15	1.184 ± 0.002	1.145	3.4	85-lav/jak, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.080	0.8	71-ric/rog <sup>c</sup>			1.174	0.9	85-tam/mur, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.096	-0.6	75-bur/ric <sup>c</sup>			1.169	1.3	90-sek/ven <sup>c</sup>
		1.069	1.9	85-mar/bha, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.220	-3.0	95-osw/pat <sup>c</sup>
		1.076	1.2	85-tam/mur, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	308.15	1.236 ± 0.001	1.219	1.4	80-aic/tar, <sup>c</sup> 81-aic/tar, <sup>c</sup> 81-aic/tar-1 <sup>c</sup>
298.15	1.135 ± 0.002	1.130	0.4	61-shi/hil, <sup>c</sup> 72-ewi/mar, <sup>c</sup> 78-kiy/hal <sup>c</sup>			1.218	1.5	83-nat/tri, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.155	-1.7	74-jai/nor, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	313.15	1.291 ± 0.001	1.301	-0.8	63-brz/har <sup>c</sup>
		1.142	-0.6	75-bur/ric <sup>c</sup>	318.15	1.348 ± 0.001	1.331	1.3	80-aic/tar, <sup>c</sup> 81-aic/tar, <sup>c</sup> 81-aic/tar-1 <sup>c</sup>
		1.126	0.8	78-gro/wil <sup>c</sup>	318.35	1.351 ± 0.001	1.335	1.2	85-lav/jak, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.120	1.3	80-aic/tar, <sup>c</sup> 81-aic/tar, <sup>c</sup> 81-aic/tar-1 <sup>c</sup>	323.15	1.410 ± 0.001	1.414	-0.3	63-brz/har <sup>c</sup>
		1.124	1.0	83-nat/tri, <sup>d</sup> 85-tam/mur, <sup>d</sup>	333.15	1.543 ± 0.001	1.540	0.2	63-brz/har <sup>c</sup>
				95-fuj/tam, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.520	1.5	80-aic/tar, <sup>c</sup> 81-aic/tar, <sup>c</sup> 81-aic/tar-1 <sup>c</sup>
		1.128	0.6	85-cos/bha <sup>c</sup>	337.95	1.613 ± 0.002	1.573	2.5	85-lav/jak, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.120	1.3	85-mar/bha, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	343.15	1.693 ± 0.002	1.679	0.8	63-brz/har <sup>c</sup>
		1.129	0.5	87-oho/tam, <sup>c</sup> 83-tam/oho, <sup>c</sup>	352.05	1.844 ± 0.002	1.885	-2.2	85-lav/jak, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
				97-oho/tam <sup>c</sup>					



Table 5. (Continued)

TK	$\beta_T/\text{GPa}^{-1}$		$\delta\beta_T/\%$ <sup>b</sup>	ref(s)	TK	$\beta_T/\text{GPa}^{-1}$		$\delta\beta_T/\%$ <sup>b</sup>	ref(s)
	eq 1 <sup>a</sup>	lit.				eq 1 <sup>a</sup>	lit.		
Methylcyclohexane									
293.15	1.082 ± 0.005	1.108	-2.3	85-tam/mur, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.126 ± 0.005	1.181	-4.6	97-oho/tam <sup>c</sup>
298.15	1.126 ± 0.005	1.16	-2.9	61-shi/hil <sup>c</sup>	303.15	1.172 ± 0.005	1.198	-2.2	74-rao/nai, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.152	-2.3	85-tam/mur, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>			1.197	-2.1	85-tam/mur, <sup>d</sup> 91-trc, <sup>e</sup> 96-zab/ruz <sup>f</sup>
		1.175	-4.2	87-oho/tam <sup>c</sup>					
Cyclooctane									
298.15	0.798 <sup>f</sup>	0.803	-0.6	74-ewi/mar, <sup>c</sup> 77-ewi/mar <sup>c</sup>	303.00	0.822 <sup>f</sup>	0.829	-0.8	74-jai/nor, <sup>d</sup> 70-ewi/lev, <sup>e</sup> 74-ewi/mar, <sup>e</sup> 96-zab/ruz <sup>f</sup>
<i>trans</i> -Bicyclo[4.4.0]decane									
298.15	0.735 ± 0.001	0.724	1.5	85-let/bax <sup>c</sup>	298.15	0.735 ± 0.001	0.764	-3.8	89-ohn/fuj <sup>c</sup>
Bicyclohexyl									
298.15	0.65 ± 0.02	0.674	-3.6	86-tar/dia <sup>c</sup>					
1,3-Cyclohexadiene									
293.15	0.962 ± 0.002	0.968	-0.6	75-bur/ric <sup>c</sup>	298.15	1.007 ± 0.002	0.962	4.6	86-tar/dia <sup>c</sup>
298.15	1.007 ± 0.002	1.012	-0.5	75-bur/ric <sup>c</sup>					
1,4-Cyclohexadiene									
293.15	0.853 ± 0.002	0.862	-1.0	75-bur/ric <sup>c</sup>	298.15	0.908 ± 0.002	0.944	-3.8	86-tar/dia <sup>c</sup>
298.15	0.908 ± 0.002	0.917	-1.0	75-bur/ric <sup>c</sup>					
Cyclohexene									
293.15	0.993 ± 0.016	1.010	-1.7	44-sch, <sup>d</sup> 72-ker/apa, <sup>e</sup> 96-zab/ruz <sup>f</sup>	298.15	1.043 ± 0.016	1.036	-0.7	86-tar/dia <sup>c</sup>
		0.983	-1.0	75-bur/ric <sup>c</sup>	303.15	1.096 ± 0.018	1.096	-0.0	49-wei, <sup>d</sup> 72-ker/apa, <sup>e</sup> 96-zab/ruz <sup>f</sup>
298.15	1.043 ± 0.016	1.041	-0.2	75-bur/ric <sup>c</sup>			1.077	-1.8	90-sek/ven <sup>d,e</sup> 96-zab/ruz <sup>f</sup>

<sup>a</sup> Uncertainty is estimated as  $\pm 2s$ , where  $s$  is a standard deviation derived from a covariance matrix of each fit. <sup>b</sup>  $[\beta_T(\text{eq 1}) - \beta_T(\text{lit})] \cdot 100 / \beta_T(\text{lit})$ . <sup>c</sup> Isothermal compressibility,  $\beta_T = -(1/V)(\partial V/\partial P)_T$ . <sup>d</sup> Sound speed. <sup>e</sup> Density and thermal expansivity,  $\alpha_P = (1/V)(\partial V/\partial T)_P$ . <sup>f</sup> Isobaric heat capacity. <sup>g</sup> Narrow temperature range fit (see Table 4). <sup>h</sup> Literature values calculated from smoothing equation representing data [67-orw/flo]. <sup>i</sup> Extrapolated beyond the temperature range of the fit (Table 3).

jac], who selected the  $P$ - $\rho$ - $T$  data measured by Grigoriev's group [75-gri/mur, 75-ras/gri] in deriving the fundamental equation (Helmholz energy) for cyclohexane based on various experimental property data.

The data available for methylcyclohexane are not in good mutual agreement. The selection of the data sets retained for the final fit and setting statistical weights was influenced by comparisons of isothermal compressibilities calculated from repeated fits with literature data. The final fit gives compressibilities on average 2.9% lower than available values in the range 293–303 K.

One or two data sets were available for other cycloalkanes. For some of them isothermal compressibility data were available for comparison (cyclooctane, *trans*-bicyclo[4.4.0]decane, bicyclohexyl) and deviations below 4% were observed (Table 5). It should be mentioned that the accuracy of the density data presented by Voss and Sloan [89-vos/slo] (cyclohexane, ethylcyclohexane, *cis*-1,2-dimethylcyclohexane) is declared by the authors to be unusually low, being 2%.

**Cycloalkenes.** The fits for 1,3-cyclohexadiene and 1,4-cyclohexadiene are recorrelations of F-type data calculated from equations presented by Burkat and Richard [75-bur/ric]. Naturally, the agreement of calculated isothermal compressibilities with values from the same source is within 1%. Larger deviations are observed for the literature values of Tardajos et al. [86-tar/dia].

The data of Melikhov et al. [91-mel/mel] for cyclohexene agree with those of Burkat and Richard [75-bur/ric] slightly better than the rejected data of Kerimov and Apaev [72-ker/apa] (see also cyclohexane); however, the average deviation of the data of Melikhov et al. [91-mel/mel] from the fit is over 0.3% (Table 4). The average deviation in the isothermal compressibility in the range 293–303 K is 0.9%.

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## Appendix: Representation of Reference Density Data $\rho[T, P_{\text{ref}}(T)] = \rho(T)$

The functions of temperature selected to represent reference density data  $\rho[T, P_{\text{ref}} = 0.101325 \text{ MPa}$  or  $P_{\text{ref}} = P_{\text{sat}}(T)] = \rho(T)$  are as follows

$$\rho(T/\text{K})/(\text{kg}\cdot\text{m}^{-3}) = \rho_c \left\{ 1 + \sum_{i=0}^5 a_i (1 - T_r)^{(i+1)/3} \right\}, \quad T_r = T/T_c \quad (\text{A1})$$

$$\rho(T/\text{K})/(\text{kg}\cdot\text{m}^{-3}) = \sum_{i=0}^2 a_i (T/100)^i \quad (\text{A2})$$

The values of the adjustable parameters  $a_i$  obtained by fitting to selected data using a weighted least-squares method are recorded in Table 6 along with some characteristics of the fits. The results in Table 6 are mostly the fits of recommended values available in *Thermodynamics Research Center Thermodynamic Tables—Hydrocarbons* (TRC Tables) combined in several cases with data either presented by authors of density data at elevated pressures or obtained by extrapolation to the reference pressure. The statistical weights of density values taken from TRC Tables were related to the number of significant digits presented in the TRC Tables; in other cases the weights were estimated from information in original papers or related to the average deviations of the extrapolations. The absence of extremes and inflection points on the function  $\rho(T)$  was checked for all fits. Equation A1 was preferably used in those cases where respective critical parameters ( $\rho_c$ ,  $T_c$ ) were available and the data covered a wide temperature range to enlarge the use of the fits beyond the temperature range of the data, particularly in the region between  $T_{\text{max}}$  and  $T_c$ . Therefore, the fits by eq A1 with  $a_0 > 0$  were

**Table 6. Parameters  $a_i$  of Smoothing Functions A1 or A2 Fitted to Selected Density Values,  $\rho[T, P=0.1 \text{ MPa or } P_{\text{sat}}(T)]$ , Critical Densities<sup>a</sup> ( $\rho_c$ ), Critical Temperatures<sup>a</sup> ( $T_c$ ), Temperature Ranges of Density Data ( $T_{\text{min}}$  and  $T_{\text{max}}$ ), and RMSD of the Fits**

eq	$a_0$	$a_1$	$a_2$	$a_3$	$a_4$	$a_5$	$\rho_c^j$ kg·m <sup>-3</sup>	$T_c$ /K	$T_{\text{min}}$ /K	$T_{\text{max}}$ /K	RMSD/ kg·m <sup>-3</sup>	ref
A1	2.019558	-1.592068	6.829264	-8.866790	4.405885		235.785	460.43	113.26	443.15	0.322	96-trc
A1	1.146314	3.273773	-4.781257	3.136329			238.120	433.78	256.58	413.15	0.123	96-trc
A1	5.722272	-28.29310	78.058688	-90.89560	39.000958		234.815	497.50	273.15	473.15	0.425	92-trc
A1	4.031489	-11.82482	24.255863	-18.32954	4.334045		234.815	504.50	293.15	473.15	0.703	92-trc
A1	1.514059	2.294405	-3.185144	2.087656			240.048	488.78	263.15	473.15	0.655	92-trc
A1	1.515081	2.028384	-2.293512	1.430761			240.718	499.98	207.93	463.15	0.273	93-bao/cac, 92-trc
A1	34.474530	-176.7741	359.19912	-321.2713	107.54832		240.875	540.64	273.15	413.15	0.168	91-trc-1
A1	2.111867	1.260332	-3.513467	3.029976			240.875	520.50	273.15	353.15	0.011	91-trc-1
A1	5.668622	-9.498036	6.609025				251.769	531.17	298.15	353.15	0.686	91-trc-1
A1	1.983676	-1.937998	10.308726	-17.09974	12.135390	-2.5519	244.083	543.96	173.15	541.15	0.254	90-trc
A1	0.973342	3.647183	-4.409424	2.744414			246.134	735.00	293.15	613.15	0.165	73-trc
A1	16.746836	-88.91311	201.73131	-201.5217	75.498293		246.848	746.00	301.32	573.15	0.096	73-trc
A2	986.90051	-67.95142							310.93	408.15	0.066	59-low/spe
A1	229.21878	-844.2745	1044.9267	-429.8758			247.420	768.00	309.75	423.15	0.181	73-trc
A1	0.564648	3.006777	-0.800313				247.420	768.00	373.15	573.15	0.450	64-doo
A2	984.53036	-67.26151							323.08	368.26	0.712	88-pet/spi <sup>d</sup>
A2	985.79012	-64.68799							333.52	371.22	0.279	87-pet/van <sup>d</sup>
A2	993.03187	-65.12813							310.93	408.15	0.036	58-cut/mcm, 59-low/spe
A1	0.668899	2.176314					249.747	857.00	373.15	573.15	0.548	64-doo
A2	991.66650	-62.70545							298.15	353.15	0.022	70-kus/tas
A2	998.52323	-63.70262							310.93	408.15	0.067	59-low/spe
A2	1000.9301	-62.17320							310.93	408.15	0.068	59-low/spe
A2	969.12134	-46.83596	-1.614357						423.15	573.15	0.062	64-doo
A2	926.32880	-66.12200	-6.000000						273.15	323.15	0.000	95-trc-2
A1	2.685307	-6.022579	18.874341	-23.26466	10.842690		237.744	464.78	183.15	448.15	0.435	86-trc-1, 51-day/fel <sup>d</sup>
A2	865.37793	-43.70657	-9.310863						203.15	353.15	0.402	86-trc-1
A2	993.20963	-100.020							293.15	298.15	0.000 <sup>f</sup>	75-bur/ric
A2	989.29900	-92.000							293.15	298.15	0.000 <sup>f</sup>	75-bur/ric
A2	887.62113	-67.020							293.15	298.15	0.000 <sup>f</sup>	75-bur/ric
A1	3.997552	-14.24428	37.259271	-40.50644	16.509841		240.461	504.03	153.15	473.15	0.322	95-trc
A1	4.694761	-10.48309	14.073957	-5.554842			243.040	537.29	223.15	523.15	0.741	95-trc-1, 81-gus/naz <sup>d</sup>

Table 6. Continued

eq	$a_0$	$a_1$	$a_2$	$a_3$	$a_4$	$a_5$	$\rho_c^j$ kg·m <sup>-3</sup>	$T_c$ /K	$T_{\min}$ /K	$T_{\max}$ /K	RMSD/ kg·m <sup>-3</sup>	ref
A1	7.845432	-20.70510	24.728860	-9.030274			243.946	566.70	273.15	393.15	0.061	86-trc
A2	960.14191	-78.76730	-0.002380						198.00	393.15	0.659	84-gus/gal, 86-trc
A2	1050.7355	-66.74182							298.15	353.15	0.062	70-kus/tas
A1	8.142771	-24.21655	31.749720	-13.20848			275.037	511.70	188.63	353.15	0.174	92-bao/cac, 95-trc-3
A1	10.138970	-51.57038	121.91233	-125.5587	48.396848		263.829	532.73	253.15	393.15	0.081	95-trc-3
A1	1.626418	0.995043	-0.255766	-0.087998	0.474534		273.251	553.50	273.15	553.15	0.179	91-trc
A2	1182.0050	-87.63877	-3.786667						288.15	313.15	0.000 <sup>i</sup>	84-mat/van
A1	1.626075	1.195703	-0.691798	1.908529	-4.062831	2.9225	266.816	572.20	143.15	568.15	0.168	91-trc
A2	1062.0709	-85.68000							298.15	353.15	0.024	70-kus/tas
A1	4.074781	-5.478427	4.241512				269.101	602.40	263.15	383.15	0.236	92-trc-1
A1	2.085256	-0.339364	0.991098				269.101	601.00	263.15	343.15	0.159	91-trc
A1	1.258428	1.535662					273.696	647.20	293.15	393.85	0.314	55-kus, 78-gou <sup>d</sup>
A1	1.523238	1.203265					280.432	687.10	298.15	353.15	0.197	70-kus/tas
A2	1050.0535	-96.32193	3.666554						253.15	393.15	0.086	92-trc-1
A2	1059.0066	-77.32000	5.000000						298.20	338.20	0.000 <sup>i</sup>	88-sid/tej
A2	1159.4907	-61.15625							310.93	408.15	0.081	59-low/spe
A2	1079.22782	-64.84824							310.95	408.15	0.147	58-cut/mcm
A2	1014.2506	-40.66359	-3.233775						333.15	408.15	0.132	58-cut/mcm
A2	1065.4843	-73.66552	1.091535						310.95	408.15	0.123	58-cut/mcm
A2	1053.1538	-63.79427							310.95	408.15	0.122	58-cut/mcm
A2	1026.5511	-65.87219							310.95	408.15	0.301	58-cut/mcm
A2	1021.2846	-65.50544							310.95	408.15	0.150	58-cut/mcm
A2	1111.3631	-61.42616							310.93	408.15	0.033	59-low/spe
A2	1623.1509	-266.9800							293.15	298.15	0.000 <sup>f</sup>	75-bur/ric
A2	2270.3436	-482.0200							293.15	298.15	0.000 <sup>f</sup>	75-bur/ric
A1	1.328770	1.273155					288.028	560.48	293.15	373.15	0.073	91-mel/mel <sup>d</sup>

<sup>a</sup> From database 93-cda. Critical densities are given with three decimal points, since they were calculated from rounded values of critical molar volumes recorded in the database 93-cda. <sup>b</sup> Combination of data in temperature ranges: 207.93–298.17 K [93-bao/cac]; 283.15–463.15 K [92-trc]. <sup>c</sup> Average deviation between two fits in the interval 373–423 K is about 1.1 kg·m<sup>-3</sup>. <sup>d</sup> Fit of values obtained by extrapolation along an isotherm of elevated-pressure data to atmospheric or saturation pressure ( $P_{\text{ref}}$ ) using the Tait equation. <sup>e</sup> Combination of data in temperature ranges: 183.15–323.15 K [86-trc-1]; 353.15–448.15 K [51-day/fel]. <sup>f</sup> Two density values available. <sup>g</sup> Combination of data in temperature ranges: 223.15–363.15 K [95-trc-1]; 373.15–523.15 K [81-gus/naz]. <sup>h</sup> Combination of data in temperature ranges: 198.00–248.00 K [84-gus/gal]; 273.15–393.15 K [86-trc]. <sup>i</sup> Combination of data in temperature ranges: 188.63–302.08 K [92-bao/cac]; 263.15–353.15 K [95-trc-3]. <sup>j</sup> Polynomial interpolation. <sup>k</sup> Combination of data in temperature ranges: 293.15–353.15 K [55-kus]; 373.75–393.85 K [78-gou].

selected to get the correct derivative of the liquid–vapor saturation line at the critical point,  $(d\rho/dT) \rightarrow -\infty$  for  $T \rightarrow T_c$ .

In the cases of the  $P$ – $\rho$ – $T$  data sets for 1-octene and cyclohexane (denoted by the letter e in the last column of Table 4) where the reference density values  $\rho(T, P_{\text{ref}}(T))$  (see

eq 1) were not available (nor original values, nor extrapolated ones), the reference density values calculated from the fits presented in Table 6 were employed in fitting the  $P$ - $\rho$ - $T$  data by the Tait equation. The fits for other substances are presented in Table 6 as auxiliary information here, which may be useful particularly for less common substances. The functions can be used to calculate smoothed reference density values for evaluation of compressed liquid densities  $\rho(T,P)$  from eq 1. Except for the fits of the values taken from the TRC Tables, the fits in Table 6 do not represent critically selected experimental data.

The densities [92-trc] for 2,3-dimethylbutane at low temperatures do not agree with the data of Baonza et al. [93-bao/cac] and Riembauer et al. [90-rie/sch]. Besides that, the fit of the values of the TRC Tables [92-trc] required seven parameters in eq A1 to get a satisfactory result. Replacement of the low-to-ambient-temperature densities [92-trc] by the values reported by Baonza et al. [93-bao/cac] led to the fit with four parameters. Two fits are presented in Table 6 for eicosane; the deviations of the first fit from the second one in the overlapping temperature range are rather large (from  $-1.1 \text{ kg}\cdot\text{m}^{-3}$  to  $2.6 \text{ kg}\cdot\text{m}^{-3}$ ). The densities of 1-heptene at saturation pressure [83-mcg] obtained by the extrapolation of data [81-gus/naz] are consistent with the values of the TRC Tables [95-trc-1]; the fit of the combined data in Table 6 exhibits, however, the point of inflection at  $T = 517 \text{ K}$ . The fit for cyclopentane does not cover the entire range of data [95-trc-3], since large deviations were observed for  $T > 363 \text{ K}$  (up to the highest temperature  $393.15 \text{ K}$ ); the fit could be improved by enlargement of the number of parameters in eq A1, but the fits with  $a_0 < 0$  and two points of inflection resulted. Extrapolated values [78-gou] for cycloheptane are by about  $1.5 \text{ kg}\cdot\text{m}^{-3}$  lower than the direct data of Kuss and Taslimi [70-kus/tas]; therefore, the two data sets were not combined for the fit. The data for cyclooctane from the same laboratory [55-kus] were combined with extrapolated values [78-gou] for  $T \geq 373.15 \text{ K}$ ; the values for lower temperatures were rejected, since they differ from the data of Kuss [55-kus] by  $-1.5 \text{ kg}\cdot\text{m}^{-3}$ .

## Literature Cited

- 29-fre/hub Fryer, E. B.; Hubbard, J. C.; Andrews, D. H. Sonic Studies of the Physical Properties of Liquids. I. The Sonic Interferometer. The Velocity of Sound in Some Organic Liquids and their Compressibilities. *J. Am. Chem. Soc.* **1929**, *51*, 759–770.
- 31-bri Bridgman, P. W. The Volume of Eighteen Liquids as a Function of Pressure and Temperature. *Proc. Am. Acad. Arts Sci.* **1931**, *66*, 185–233.
- 32-bri Bridgman, P. W. Volume–Temperature–Pressure Relations for Several Nonvolatile Liquids. *Proc. Am. Acad. Arts Sci.* **1932**, *67*, 1–27.
- 34-rot/nag Rotinjan, L.; Nagornow, N. Die Zustandsflächen des Cyclohexans. (Phase Surfaces of Cyclohexane). *Z. Phys. Chem., Abt. A* **1934**, *169*, 20–30.
- 40-ke/fel Kelso, E. A.; Felsing, W. A. The Pressure–Volume–Temperature Relations of *n*-Hexane and of 2-Methylpentane. *J. Am. Chem. Soc.* **1940**, *62*, 3132–3134.
- 40-sch/hof Schoch, E. P.; Hoffmann, A. E.; Mayfield, F. D. Solubility of Methane in Cyclohexane. *Ind. Eng. Chem.* **1940**, *32*, 1351–1353.
- 42-ke/fel Kelso, E. A.; Felsing, W. A.  $P$ - $V$ - $T$  Relations and Derived Quantities for Hexanes. *Ind. Eng. Chem.* **1942**, *34*, 161–163.
- 43-fel/wat Felsing, W. A.; Watson, G. M. The Pressure–Volume–Temperature Relations of 2,2-Dimethylbutane. *J. Am. Chem. Soc.* **1943**, *65*, 1889–1891.
- 43-fel/wat-1 Felsing, W. A.; Watson, G. M. The Pressure–Volume–Temperature Relations of 2,2,4-Trimethylpentane. *J. Am. Chem. Soc.* **1943**, *65*, 780–781.
- 44-sch Schaaffs, W. Untersuchungen über Schallgeschwindigkeit und Konstitution. I. Teil: Die Schallgeschwindigkeit in organischen Flüssigkeiten. (A Study of Speed of Sound and Constitution. I. Speed of Sound in Organic Liquids.) *Z. Phys. Chem. (Leipzig)* **1944**, *194*, 28–38.
- 48-lag/mcm Lagemann, R.; McMillan, D.; Woolsey, M. Ultrasonic Velocity in Series of 1-Olefins. *J. Chem. Phys.* **1948**, *16*, 247–249.
- 49-bri Bridgman, P. W. Further Rough Compressions to 40,000  $\text{kg}/\text{cm}^2$ , Especially Certain Liquids. *Proc. Am. Acad. Arts Sci.* **1949**, *77*, 129–146.
- 49-wei Weissler, A. Ultrasonic Investigation of Molecular Properties of Liquids. IV. Cyclic Compounds. *J. Am. Chem. Soc.* **1949**, *71*, 419–421.
- 51-day/fel Day, H. O.; Felsing, W. A. The Compressibility of Pentene-1. *J. Am. Chem. Soc.* **1951**, *73*, 4839–4840.
- 52-day/fel Day, H. O.; Felsing, W. A. The Pressure–Volume–Temperature Relations of 3-Methylpentane. *J. Am. Chem. Soc.* **1952**, *74*, 1951–1953.
- 52-jac Jacobson, B. Intermolecular Free Lengths in the Liquid State. I. Adiabatic and Isothermal Compressibilities. *Acta Chem. Scand.* **1952**, *6*, 1485–1498.
- 54-isa/li Isaac, R.; Li, K.; Canjar, L. N. Volumetric Behavior of Isopentane. *Ind. Eng. Chem.* **1954**, *46*, 199–201.
- 55-kus Kuss, E. Hochdruckuntersuchungen III: Die Viskosität von komprimierten Flüssigkeiten. (High-Pressure Investigation III: Viscosity of Compressed Liquids). *Z. Angew. Phys.* **1955**, *7*, 372–378.
- 57-rea/sag Reamer, H. H.; Sage, B. H. Phase Equilibria in Hydrocarbon Systems. Volumetric Behavior of Cyclohexane. *Chem. Eng. Data Ser.* **1957**, *2*, 9–12.
- 58-cut/mcm Cutler, W. G.; McMickle, R. H.; Webb, W.; Schiessler, R. W. Study of the Compressions of Several High Molecular Weight Hydrocarbons. *J. Chem. Phys.* **1958**, *29*, 727–740.
- 58-par/pan Parthasarathy, S.; Pancholy, M.; Chhapgar, A. F. Ultrasonic Absorption in Some Homologous Series of Organic Liquids. Part II. Hydrocarbons. *Nuovo Cimento* **1958**, *10*, 118–131.
- 59-gol/vag Golubev, I. F.; Vagina, E. N. Specific Mass of Benzene, Cyclohexane, and their Mixtures at High Pressures and Various Temperatures. *Tr. GIAP* **1959**, *9*, 95–107 (in Russian).
- 59-low/spe Lowitz, D. A.; Spencer, J. W.; Webb, W.; Schiessler, R. W. Temperature–Pressure–Structure Effects on the Viscosity of Several Higher Hydrocarbons. *J. Chem. Phys.* **1959**, *30*, 73–83.
- 61-shi/hil Shinoda, K.; Hildebrand, J. H. Compressibilities and Isochores of  $(\text{C}_3\text{H}_7\text{COOCH}_2)_4\text{C}$ ,  $\text{c-Si}_4\text{O}_4(\text{CH}_3)_8$ ,  $n\text{-C}_5\text{H}_{12}$ ,  $n\text{-C}_8\text{H}_{18}$ ,  $2,2,4\text{-C}_5\text{H}_9(\text{CH}_3)_3$ ,  $\text{c-C}_5\text{H}_{10}$ ,  $\text{c-C}_6\text{H}_{12}$ ,  $\text{c-C}_6\text{H}_{11}\text{CH}_3$ ,  $\text{C}_6\text{H}_5\text{CH}_3$ ,  $p\text{-C}_6\text{H}_4(\text{CH}_3)_2$ ,  $s\text{-C}_6\text{H}_3(\text{CH}_3)_3$ ,  $\text{CH}_2\text{Cl}_2$ . *J. Chem. Phys.* **1961**, *65*, 183–183.
- 62-hol/wha Holder, G. A.; Whalley, E. Compressibility of Liquids. Part 1. Experimental Methods, and Compressibility of Carbon Tetrachloride, Benzene and Cyclohexane. *Trans. Faraday Soc.* **1962**, *58*, 2095–2107.
- 63-brz/har Brzostowski, W.; Hardman, T. M. Thermodynamics of Cyclohexane–Propanol Mixtures. I. Volume Properties. *Bull. Acad. Pol. Sci., Ser. Sci. Chim.* **1963**, *11*, 447–452.
- 64-doo Doolittle, A. K. Specific Volumes of *n*-Alkanes. *J. Chem. Eng. Data* **1964**, *9*, 275–279.
- 67-orw/flo Orwoll, R. A.; Flory, P. J. Equation-of-State Parameters for Normal Alkanes. Correlation with Chain Length. *J. Am. Chem. Soc.* **1967**, *89*, 6814–6822.
- 68-gon/lee Gonzales, M. H.; Lee, A. L. Viscosity of 2,2-Dimethylpropane. *J. Chem. Eng. Data* **1968**, *13*, 66–69.
- 69-bra/fre Brazier, D. W.; Freeman, G. R. The Effect of Pressure on the Density, Dielectric Constant, and Viscosity of Several Hydrocarbons and Other Organic Liquids. *Can. J. Chem.* **1969**, *47*, 893–899.

- 69-mop Mopsik, F. I. Dielectric Properties of Slightly Polar Organic Liquids as a Function of Pressure, Volume, and Temperature. *J. Chem. Phys.* **1969**, *50*, 2559–2569.
- 70-aba/ker Abas-Zade, A. K.; Kerimov, A. M.; Agaev, N. A.; Apaev, T. A. Experimental Determination of Density of Hydrocarbons. *Teplofiz. Svoistva Zhidk., Nauka, Moskva* **1970**, 34–38 (in Russian).
- 70-ewi/lev Ewing, M. B.; Levien, B. J.; Marsh, K. N.; Stokes, R. H. Excess Enthalpies, Excess Volumes, and Excess Gibbs Free Energies for Mixtures of Cyclooctane + Cyclopentane at 288.15, 298.15, and 308.15 K. *J. Chem. Thermodyn.* **1970**, *2*, 689–695.
- 70-kus/tas Kuss, E.; Taslimi, M. p, V, T-Messungen an Zwanzig Organischen Flüssigkeiten (p, V, T Measurement of Twenty Organic Liquids). *Chem.-Ing.-Technol.* **1970**, *42*, 1073–1081.
- 71-hou/hey Houck, J. C.; Heydemann, P. L. M. Combined Low-Pressure and High-Pressure Measurements of Density and Bulk Modulus of Aviation Instrument Oil and 2-Methylbutane and their Mixtures. *J. Res. Natl. Bur. Stand., Sect. A* **1971**, *75*, 121–127.
- 71-ric/rog Richard, A. J.; Rogers, K. S. The Isothermal Compressibility of Organic Liquids by Ultracentrifugation. Correlation with Surface Tension. *Can. J. Chem.* **1971**, *49*, 3956–3959.
- 72-ewi/mar Ewing, M. B.; Marsh, K. N.; Stokes, R. H. A Dilution Piezometer for Isothermal Compressibilities of Mixtures. Excess Compressibilities of Benzene + Cyclohexane at 298.15 K. *J. Chem. Thermodyn.* **1972**, *4*, 637–646.
- 72-gol/ada Golik, A. Z.; Adamenko, I. I.; Borovik, V. V. Observation of P–V–T Relation for n-Paraffins in the Interval of Pressure up to 2500 atm and Temperatures up to 120 degrees Centigrade. *Ukr. Fiz. Zh. (Ukr. Ed.)* **1972**, *17*, 2075–2078 (in Russian).
- 72-ker/apa Kerimov, A. M.; Apaev, T. A. Experimental Values of Density of Hexene-1, Octene-1, Cyclohexene, Cyclohexane, and Methylcyclohexane in Dependence on Temperature and Pressure. *Teplofiz. Svoistva Veshchestv Mater.* **1972**, *5*, 26–46 (in Russian).
- 72-lys Lysne, P. C. Nonlinear U(u) Hugoniot of Liquids at Low Pressures. *J. Chem. Phys.* **1972**, *57*, 492–494.
- 73-daw/sil Dawson, P. P.; Silberberg, I. H.; McKetta, J. J. Volumetric Behavior, Vapor Pressures, and Critical Properties of Neopentane. *J. Chem. Eng. Data* **1973**, *18*, 7–51.
- 73-ker/apa Kerimov, A. M.; Apaev, T. A. Experimental Observation of P–V–T Dependence of Cyclohexane in a Wide Range of Temperatures and Pressures. *Teplofiz. Svoistva Zhidk., Nauka, Moskva* **1973**, 84–88 (in Russian).
- 73-koh/luk Kohn, J. P.; Luks, K. D. Isothermal Compressibility of Liquids at Low Pressures. *Chem. Eng. Commun.* **1973**, *1*, 107–109.
- 73-rog/bur Rogers, K. S.; Burkat, R.; Richard, A. J. The Ultracentrifuge as a Pressure-Densitometer. *Can. J. Chem.* **1973**, *51*, 1183–1186.
- 73-trc TRC Tables 23-2-(1.101)-d. C–H. Normal Alkanes (Paraffins), C<sub>1</sub> to C<sub>20</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1973; pp 1010–1015.
- 74-apa/ker Apaev, T. A.; Kerimov, A. M. Experimental Observation of Density of Cyclohexane at High Pressures and Various Temperatures. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1974**, *17* (4), 78–106 (in Russian).
- 74-ewi/mar Ewing, M. B.; Marsh, K. N. Excess Functions for Cyclopentane + Cyclohexane, Cyclopentane + Cycloheptane, and Cyclohexane + Cyclo-octane. *J. Chem. Thermodyn.* **1974**, *6*, 395–406.
- 74-hou Houck, J. C. High-Pressure Measurements of Density, Velocity of Sound, and Bulk Moduli of Pentane and 2-Methylbutane and their Mixtures. *J. Res. Natl. Bur. Stand., Sect. A* **1974**, *78*, 617–622.
- 74-jai/nor Jain, D. V. S.; North, A. M.; Pethrick, R. A. Adiabatic Compressibility of Binary Liquid Mixtures. *J. Chem. Soc., Faraday Trans. 1* **1974**, *70*, 1292–1298.
- 74-rao/nai Rao, M. V. P.; Naidu, P. R. Isentropic Compressibilities of Mixtures of an Alcohol + Methylcyclohexane. *J. Chem. Thermodyn.* **1974**, *6*, 1195–1196.
- 75-bur/ric Burkat, R. K.; Richard, A. J. Low-pressure Studies of the Isothermal Compressibilities and Specific Volumes of Organic Liquids. *J. Chem. Thermodyn.* **1975**, *7*, 271–277.
- 75-gri/mur Grigoriev, B. A.; Murdaev, R. M.; Rastorguev, Yu. L. Experimental Observation of P–V–T Dependence of Cyclohexane at High Temperatures and Pressures. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1975**, *18* (3), 61–63 (in Russian).
- 75-luk/dav Luks, K. D.; Davis, H. T.; Kohn, J. P. The Isothermal Compressibility of the Liquid Mixtures Neopentane + n-Hexane and n-Hexane + n-Decane at Low Pressures and 298.15 K: Experiment and Theory. *J. Chem. Thermodyn.* **1975**, *7*, 311–318.
- 75-ras/gri Rastorguev, Yu. L.; Grigoriev, B. A.; Murdaev, R. M. Experimental Observation of P–V–T Dependence of Cyclohexane in Liquid Phase. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1975**, *18* (1), 66–70 (in Russian).
- 75-zor/hen Zordan, T. A.; Henry, R. M. Volumetric Properties of Liquid Propylene. *J. Chem. Eng. Data* **1975**, *20*, 343–347.
- 76-sah/gag Sahli, B. P.; Gager, H.; Richard, A. J. Ultracentrifugal Studies of the Isothermal Compressibilities of Organic Alcohols and Alkanes. Correlation with Surface Tension. *J. Chem. Thermodyn.* **1976**, *8*, 179–188.
- 77-ewi/mar Ewing, M. B.; Marsh, K. N. Isothermal Compressibilities of Cyclopentane + Cyclo-octane and + Octamethylcyclotetrasiloxane at 298.15 K. *J. Chem. Thermodyn.* **1977**, *9*, 371–374.
- 78-gou Gouel, P. Density of Alkanes (C<sub>6</sub> to C<sub>16</sub>), Cycloalkanes, and Alkylbenzenes. *Bull. Cent. Rech. Explor.-Prod. Elf-Aquitaine* **1978**, *2*, 211–225 (in French).
- 78-gro/wil Grolier, J.-P. E.; Wilhelm, E.; Hamed, M. H. Molar Heat Capacity and Isothermal Compressibility of Binary Liquid Mixtures: Carbon Tetrachloride + Benzene, Carbon Tetrachloride + Cyclohexane, and Benzene + Cyclohexane. *Ber. Bunsen-Ges. Phys. Chem.* **1978**, *82*, 1282–1290.
- 78-kiy/hal Kiyohara, O.; Halpin, C. J.; Benson, G. C. Ultrasonic Velocities, Compressibilities, and Heat Capacities for Binary Mixtures of Benzene, Cyclohexane and Tetrachloromethane at 298.15 K. *J. Chem. Thermodyn.* **1978**, *10*, 721–730.
- 79-dic Dick, R. D. Shock Compression Data for Liquids. I. Six Hydrocarbon Compounds. *J. Chem. Phys.* **1979**, *71*, 3203–3212.
- 79-dyk/rep Dykyj, J.; Repáš, M. *Vapour Pressure of Organic Compounds*; Veda: Bratislava, 1979 (in Czech).
- 79-isd/dym Isdale, J. D.; Dymond, J. H.; Brawn, T. A. Viscosity and Density of n-Hexane–Cyclohexane Mixtures Between 25 and 100 °C up to 500 MPa. *High Temp.–High Pressures* **1979**, *11*, 571–580.
- 79-jon/has Jonas, J.; Hasha, D.; Huang, S. G. Self-diffusion and Viscosity of Methylcyclohexane in the Dense Liquid Region. *J. Chem. Phys.* **1979**, *71*, 3996–4000.
- 79-kas/fuk Kashiwagi, H.; Fukunaga, T.; Tanaka, Y.; Kubota, H.; Makita, T. Dielectric Constant and Density of Cyclohexane–Benzene Mixture under High Pressure. *Rev. Phys. Chem. Jpn.* **1979**, *49*, 70–84.
- 80-aic/tar Aicart, E.; Tardajos, G.; Diaz Pena, M. Isothermal Compressibility of Cyclohexane + n-Hexane, Cyclohexane + n-Heptane, Cyclohexane + n-Octane, and Cyclohexane + n-Nonane. *J. Chem. Eng. Data* **1980**, *25*, 140–145.
- 80-jon/has Jonas, J.; Hasha, D.; Huang, S. G. Density Effects on Transport Properties in Liquid Cyclohexane. *J. Phys. Chem.* **1980**, *84*, 109–112.
- 80-oza/ooy Ozawa, S.; Ooyatsu, N.; Yamabe, M.; Honmo, S.; Ogino, Y. Specific Volumes of Binary Liquid Mixtures at High Pressures. 1. Experimental Results for (Ethanol + Methylcyclopentane), (n-Heptane + Ethanol), and (Methylcyclopentane + n-Heptane). *J. Chem. Thermodyn.* **1980**, *12*, 229–242.

- 81-aic/tar Aicart, E.; Tardajos, G.; Diaz Pena, M. Isothermal Compressibility of Cyclohexane + *n*-Tridecane and + *n*-Pentadecane at 298.15, 308.15, 318.15, and 333.15 K. *J. Chem. Thermodyn.* **1981**, *13*, 783–788.
- 81-aic/tar-1 Aicart, E.; Tardajos, G.; Diaz Pena, M. Isothermal Compressibility of Cyclohexane + *n*-Decane, Cyclohexane + *n*-Dodecane, and Cyclohexane + *n*-Tetradecane. *J. Chem. Eng. Data* **1981**, *26*, 22–26.
- 81-bue/mau Buehner, K.; Maurer, G.; Bender, E. Pressure-Enthalpy Diagrams for Methane, Ethane, Propane, Ethylene, and Propylene. *Cryogenics* **1981**, 157–164.
- 81-gus/naz Guseinov, S. O.; Naziev, Ya. M.; Shakhverdiev, A. N. Thermodynamic Properties of Heptene-1 at High Pressures. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1981**, *24* (7), 62–64 (in Russian).
- 82-bro/lop Brostow, W.; Lopez, D. M. M.; Maynadier, P. Isothermal Compressibility of Liquids: New Results on Temperature Dependence. *Proc. 8th Symp. Thermophys. Prop., Vol. 1: Thermophysical Properties of Fluids* **1982**, 122–127.
- 82-gus/gal Guseinov, S. O.; Galandarov, Z. S. Observation of Density of Hexene-1 at Low Temperatures and Various Pressures. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1982**, *25* (8), 66–68 (in Russian).
- 82-tre/han Treszczanowicz, A. J.; Handa, Y. P.; Benson, G. C. Excess Volumes and Isentropic Compressibilities of Decan-1-ol + 2,2-Dimethylbutane and + 2,2,4-Trimethylpentane. *J. Chem. Thermodyn.* **1982**, *14*, 871–881.
- 82-wis/wue Wisotzki, K. D.; Wuerflinger, A. PVT Data for Liquid and Solid Cyclohexane, Cyclohexanone and Cyclopentanol up to 3000 bar. *J. Phys. Chem. Solids* **1982**, *43*, 13–20.
- 83-aic/kum Aicart, E.; Kumaran, M. K.; Halpin, C. J.; Benson, G. C. Ultrasonic Speeds and Isentropic Compressibilities of 2-Methylpentan-1-ol with Hexane Isomers at 298.15 K. *J. Chem. Thermodyn.* **1983**, *15*, 1189–1197.
- 83-aww/pet Awwad, A. M.; Pethrick, R. A. Adiabatic Compressibility of Branched Chain Hydrocarbons—Pentanes and Hexanes. *J. Mol. Liq.* **1983**, *25*, 115–127.
- 83-gus/sha Guseinov, S. O.; Shakhverdiev, A. N.; Naziev, Ya. M. Experimental Observation of Density of Ethylcyclohexane at Various Parameters of State. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1983**, *26* (3), 50–52 (in Russian).
- 83-mcg McGarry, J. Correlation and Prediction of the Vapor Pressures of Pure Liquids over Large Pressure Ranges. *Ind. Eng. Chem., Process Des. Dev.* **1983**, *22*, 313–332.
- 83-nat/tri Nath, J.; Tripathi, A. D. Binary Systems of 1,1,2,2-Tetrachloroethane with Benzene, Toluene, *p*-Xylene, Acetone, and Cyclohexane. 1. Excess Volumes, Ultrasonic Velocities, and Adiabatic Compressibilities at 298.15 and 308.15 K. *J. Chem. Eng. Data* **1983**, *28*, 263–266.
- 83-tam/oho Tamura, K.; Ohomuro, K.; Murakami, S. Speeds of Sound, Isentropic and Isothermal Compressibilities, and Isochoric Heat Capacities of {*x*C<sub>6</sub>H<sub>12</sub> + (1-*x*)C<sub>6</sub>H<sub>6</sub>}, {*x*CCl<sub>4</sub> + (1-*x*)C<sub>6</sub>H<sub>6</sub>}, and {*x*C<sub>7</sub>H<sub>16</sub> + (1-*x*)C<sub>6</sub>H<sub>6</sub>} at 298.15 K. *J. Chem. Thermodyn.* **1983**, *15*, 859–868.
- 83-wax/gal Waxman, M.; Gallagher, J. S. Thermodynamic Properties of Isobutane from 250 to 600 K and Pressures from 0.1 to 40 MPa. *J. Chem. Eng. Data* **1983**, *28*, 224–241.
- 84-aww/pet Awwad, A. M.; Pethrick, R. A. Isentropic Compressibilities of Hydrocarbons and their Mixtures. Mixtures of Linear and Branched-chain Alkanes. *J. Chem. Thermodyn.* **1984**, *16*, 131–136.
- 84-bou/fri Boublik, T.; Fried, V.; Hála, E. *Vapour Pressure of Pure Substances*; Elsevier: Amsterdam, 1984.
- 84-gus/gal Guseinov, S. O.; Galandarov, Z. S. Observation of Density and Dynamic Viscosity of Nonene-1 at Various Temperatures and Pressures. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1984**, *27* (4), 50–55 (in Russian).
- 84-mat/van Matsuo, S.; Van Hook, A. Isothermal Compressibility of C<sub>6</sub>H<sub>6</sub>, C<sub>6</sub>D<sub>6</sub>, *c*-C<sub>6</sub>H<sub>12</sub>, *c*-C<sub>6</sub>D<sub>12</sub>, and their Mixtures from 0.1 to 35 MPa at 288, 298, and 313 K. *J. Phys. Chem.* **1984**, *88*, 1032–1040.
- 85-cos/bha Costas, M.; Bhattacharyya, S. N.; Patterson, D. Liquid Structure and the Thermal Pressure Coefficients of Cyclohexane + Normal- and Branched-Alkane Mixtures. *J. Chem. Soc., Faraday Trans. 1* **1985**, *81*, 387–395.
- 85-dym/isd Dymond, J. H.; Isdale, J. D.; Glen, N. F. Density Measurement at High Pressure. *Fluid Phase Equilib.* **1985**, *20*, 305–314.
- 85-lav/jak Lavrentjev, I. P.; Jakovlev, V. F. Speed of Ultrasound and Viscosity in Cyclohexane along the Saturation Line. *Zh. Fiz. Khim.* **1985**, *59*, 2893–2894 (in Russian).
- 85-let/bax Letcher, T. M.; Baxter, R. C. Excess Volumes and Enthalpies of Mixing Benzene with Various Bicyclic Compounds. *J. Solution Chem.* **1985**, *14*, 35–40.
- 85-mar/bha Marwein, B. L.; Bhat, S. N. Ultrasonic Study of Molecular Interactions in Ternary Liquid Systems. *Acustica* **1985**, *58*, 243–247.
- 85-tam/mur Tamura, K.; Murakami, S.; Doi, S. Speeds of Sound, Densities, and Isentropic Compressibilities of {*x*C<sub>6</sub>H<sub>12</sub> + (1-*x*)C<sub>6</sub>H<sub>5</sub>CH<sub>3</sub>}, {*x*C<sub>6</sub>H<sub>11</sub>CH<sub>3</sub> + (1-*x*)C<sub>6</sub>H<sub>6</sub>}, and {*x*C<sub>6</sub>H<sub>11</sub>CH<sub>3</sub> + (1-*x*)C<sub>6</sub>H<sub>5</sub>CH<sub>3</sub>}, from 293.15 to 303.15 K. *J. Chem. Thermodyn.* **1985**, *17*, 325–333.
- 85-tek/cib Tekáč, V.; Cibulka, I.; Holub, R. PVT Properties of Liquids and Liquid Mixtures: A Review of the Experimental Methods and the Literature Data. *Fluid Phase Equilib.* **1985**, *19*, 33–149.
- 86-hol/goe Holzapfel, K.; Goetze, G.; Kohler, F. Volume and Isothermal Compressibility of Some Normal Alkanes (C<sub>5</sub>–C<sub>16</sub>) + 2,2,4-Trimethylpentane, + Cycloalkanes (C<sub>5</sub>, C<sub>6</sub>, C<sub>8</sub>), or Methylcyclohexane. *Int. Data Ser., Sel. Data Mixtures, Ser. A* **1986**, *1*, 39–39.
- 86-jah/jac Jahangiri, M.; Jacobsen, R. T.; Stewart, R. B. Thermodynamic Properties of Ethylene from the Freezing Line to 450 K at Pressures to 260 MPa. *J. Phys. Chem. Ref. Data* **1986**, *15*, 593–734.
- 86-tar/dia Tardajos, G.; Diaz Pena, M.; Lainez, A.; Aicart, E. Speed of Sound in and Isothermal Compressibility and Isobaric Expansivity of Pure Liquids at 298.15 K. *J. Chem. Eng. Data* **1986**, *31*, 492–493.
- 86-trc TRC Tables 23-2-(5.1101)-d. C–H. 1-Alkenes (Normal Monoolefins), C<sub>2</sub> to C<sub>20</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1986; pp 2500–2502.
- 86-trc-1 TRC Tables 23-2-(5.1200)-d. C–H. Alkenes (Monoolefins), C<sub>2</sub> to C<sub>5</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1986; pp 2600–2602.
- 87-hol/goe Holzapfel, K.; Goetze, G.; Demiriz, A. M.; Kohler, F. Volume and Isothermal Compressibility of Some Normal Alkanes (C<sub>5</sub>–C<sub>16</sub>) + 2,3-Dimethylbutane, + Methylcyclopentane, + Butylcyclohexane, + Benzene, + 2-Propanone, or + Tetrachloromethane. *Int. Data Ser., Sel. Data Mixtures, Ser. A* **1987**, *1*, 30–56.
- 87-led Lederer, T. Molvolumina von reinen flüssigen Stoffen und der binären Mischung BUTAN-1-OL/N-HEXAN bei hohem Druck und verschiedenen Temperaturen. Thesis, Heidelberg University, Germany, 1987, pp 1–94.
- 87-man/cri Mansker, L. D.; Criser, A. C.; Jangkamolkulchai, A.; Luks, K. D. The Isothermal Compressibility of *n*-Paraffin Liquids at Low Pressures. *Chem. Eng. Commun.* **1987**, *57*, 87–93.
- 87-oho/tam Ohomuro, K.; Tamura, K.; Murakami, S. Speeds of Sound, Excess Molar Volumes, and Isentropic Compressibilities of {*x*CH<sub>3</sub>COC<sub>2</sub>H<sub>5</sub> + (1-*x*)C<sub>7</sub>H<sub>16</sub>}, {*x*CH<sub>3</sub>COC<sub>2</sub>H<sub>5</sub> + (1-*x*)C<sub>6</sub>H<sub>12</sub>}, {*x*CH<sub>3</sub>COC<sub>2</sub>H<sub>5</sub> + (1-*x*)C<sub>6</sub>H<sub>11</sub>CH<sub>3</sub>}, {*x*C<sub>2</sub>H<sub>5</sub>COC<sub>2</sub>H<sub>5</sub> + (1-*x*)C<sub>7</sub>H<sub>16</sub>}, and {*x*C<sub>2</sub>H<sub>5</sub>COC<sub>2</sub>H<sub>5</sub> + (1-*x*)C<sub>6</sub>H<sub>12</sub>} at 298.15 K. *J. Chem. Thermodyn.* **1987**, *19*, 163–169.

- 87-pet/van Peters, C. J.; van der Kooi, H. J.; de Swan Arons, J. Measurements and Calculations of Phase Equilibria for (Ethane + Tetracosane) and (p, Vm, T) of Liquid Tetracosane. *J. Chem. Thermodyn.* **1987**, *19*, 395–405.
- 87-sun/kor Sun, T. F.; Kortbeek, P. J.; Trappeniers, N. J.; Biswas, S. N. Acoustic and Thermodynamic Properties of Benzene and Cyclohexane as a Function of Pressure and Temperature. *Phys. Chem. Liq.* **1987**, *16*, 163–178.
- 87-you/ely Younglove, B. A.; Ely, J. F. J. Thermophysical Properties of Fluids. II. Methane, Ethane, Propane, Isobutane, and Normal Butane. *J. Phys. Chem. Ref. Data* **1987**, *16*, 577–789.
- 88-dym/mal Dymond, J. H.; Malhotra, R.; Isdale, J. D.; Glen, N. F. (p,  $\rho$ , T) of *n*-Heptane, Toluene, and Oct-1-ene in the Range 298 to 373 K and 0.1 to 400 MPa and Representation by the Tait Equation. *J. Chem. Thermodyn.* **1988**, *20*, 603–614.
- 88-mor/aon Moriyoshi, T.; Aono, T. Compressions of C6 Hydrocarbons from 298.15 to 313.15 K at Pressures to 145 MPa. *J. Chem. Thermodyn.* **1988**, *20*, 185–191.
- 88-pet/spi Peters, C. J.; Spiegelaar, J.; de Swaan Arons, J. Phase Equilibria in Binary Mixtures of Ethane + Docosane and Molar Volumes of Liquid Docosane. *Fluid Phase Equilib.* **1988**, *41*, 245–256.
- 88-sid/tej Siddiqi, S. A.; Teja, A. S. High-Pressure Densities of Mixtures of Coal Chemicals. *Chem. Eng. Commun.* **1988**, *72*, 159–169.
- 88-tre/ben Treszczanowicz, A. J.; Benson, G. C. Prediction of Excess Volumes of 1-Alkanols + Branched Aliphatic Hydrocarbon Binary Systems in Terms of an Association Model with a Flory Contribution Term. *Fluid Phase Equilib.* **1988**, *41*, 31–42.
- 89-fri/ely Friend, D. G.; Ely, J. F.; Ingham, H. Thermophysical Properties of Methane. *J. Phys. Chem. Ref. Data* **1989**, *18*, 583–638.
- 89-ohn/fuj Ohnishi, K.; Fujihara, I.; Murakami, S. Thermodynamic Properties of Decalins Mixed with Hexane Isomers at 298.15 K. II. Excess Volumes and Isentropic Compressibilities. *Fluid Phase Equilib.* **1989**, *46*, 73–84.
- 89-vos/slo Voss, S. F.; Sloan, E. D. Thermal Conductivity and Heat Capacity of Synthetic Fuel Components. *Int. J. Thermophys.* **1989**, *10*, 1029–1040.
- 90-mal/woo Malhotra, R.; Woolf, L. A. Thermodynamic Properties of 2,2,4-Trimethylpentane. *Int. J. Thermophys.* **1990**, *11*, 1059–1072.
- 90-pol/wei Polzin, B.; Weiss, A. Transport Properties of Liquids. VIII. Molar Volume and Selfdiffusion of Organic Liquids at Pressures up to 200 MPa. *Ber. Bunsen-Ges. Phys. Chem.* **1990**, *94*, 746–758.
- 90-rie/sch Riembauer, M.; Schulte, L.; Wuerflinger, A. PVT Data of Liquid and Solid Phases of Methanol, Cyclohexanol, and 2,3-Dimethylbutane up to 300 MPa. *Z. Phys. Chem. (Munich)* **1990**, *166*, 53–61.
- 90-sek/ven Sekar, P. R.; Venkateswarlu, R.; Reddy, K. S. Excess Volumes, Isentropic Compressibilities, and Viscosities of Binary Mixtures Containing Cyclohexene. *Can. J. Chem.* **1990**, *68*, 363–368.
- 90-tos/fig Toscani, S.; Figuiere, P.; Szwarc, H. Measurements of (p,  $\rho$ , T) on c-C<sub>6</sub>H<sub>12</sub> and on 0.501 c-C<sub>6</sub>H<sub>12</sub> + 0.499 n-C<sub>7</sub>H<sub>16</sub> at Pressures up to 100 MPa. *J. Chem. Thermodyn.* **1990**, *22*, 293–300.
- 90-trc TRC Tables 23-2-(1.203)-d. C–H. Alkanes (Paraffins), C<sub>8</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1990; pp 1490–1497.
- 91-fri/ing Friend, D. G.; Ingham, H.; Ely, J. F. Thermophysical Properties of Ethane. *J. Phys. Chem. Ref. Data* **1991**, *20*, 275–347.
- 91-mel/mel Melikhov, Yu. F.; Mel'nikov, G. A.; Tutov, V. M.; Verveiko, V. N. Acoustic and Thermophysical Properties of Some Derivatives of Hexane at High Parameters of State. *Izv. Vyssh. Uchebn. Zaved., Energ.* **1991**, *34* (11), 73–78 (in Russian).
- 91-pap/zia Papaioannou, D.; Ziakas, D.; Panayiotou, G. Volumetric Properties of Binary Mixtures. 1. 2-Propanone + 2,2,4-Trimethylpentane and *n*-Heptane + Ethanol Mixtures. *J. Chem. Eng. Data* **1991**, *36*, 35–39.
- 91-set/wag Setzman, U.; Wagner, W. A New Equation of State and Tables of Thermodynamic Properties for Methane Covering the Range from the Melting Line to 625 K at Pressures up to 1000 MPa. *J. Phys. Chem. Ref. Data* **1991**, *20*, 1061–1155.
- 91-tan/hos Tanaka, Y.; Hosokawa, H.; Makita, T. Viscosity and Density of Binary Mixtures of Cyclohexane with *n*-Octane, *n*-Dodecane, and *n*-Hexadecane under High Pressures. *Int. J. Thermophys.* **1991**, *12*, 245–264.
- 91-trc TRC Tables 23-2-(3.1112)-d. C–H. Alkylcyclohexanes, C<sub>6</sub> to C<sub>8</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1991; pp 2100–2101.
- 91-trc-1 TRC Tables 23-2-(1.202)-d. C–H. Alkanes (Paraffins), C<sub>7</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1991; pp 1460–1461.
- 92-bao/cac Baonza, V. G.; Caceres, M.; Nunez Delgado, J. Study of the Equation of State of Cyclopentane from 193 to 298 K and Pressures up to 104 MPa. *Ber. Bunsen-Ges. Phys. Chem.* **1992**, *96*, 1859–1868.
- 92-lag/bon Lagourette, B.; Boned, C.; Saint-Guirons, H.; Xans, P.; Zhou, H. Densimeter Calibration Method Versus Temperature and Pressure. *Meas. Sci. Technol.* **1992**, *3*, 669–703.
- 92-naz/gas Naziev, Ya. M.; Gasanov, V. G.; Naziev, D. Ya.; Allakhverdiev, A. M. Observation of P–V–T Dependence of isoctane and Octene-1 at Various Temperatures and Pressures. *Izv. Vyssh. Uchebn. Zaved., Neft Gaz* **1992**, (5), 54–56 (in Russian).
- 92-trc TRC Tables 23-2-(1.201)-d. C–H. Alkanes, C<sub>6</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1992; pp 1440–1441.
- 92-trc-1 TRC Tables 23-2-(3.1110)-d. C–H. *n*-Alkylcyclohexanes, C<sub>6</sub> to C<sub>22</sub>. *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1992; pp 2050–2051.
- 92-wal/bar Walter, T.; Bardelmeier, U.; Wuerflinger, A. Ultrasonic Studies of Rotational Isomerism and P, V, T Data of 2-Methylbutane at Elevated Pressures. *Ber. Bunsen-Ges. Phys. Chem.* **1992**, *96*, 717–721.
- 92-bao/cac Baonza, V. G.; Caceres Alonso, M.; Nunez Delgado, J. Study of the Equation of State of Liquid 2,3-Dimethylbutane at High Pressures. *J. Phys. Chem.* **1993**, *97*, 2002–2008.
- 93-cda CDATA, Database of Physical and Transport Properties of Pure Fluids. Department of Physical Chemistry, Institute of Chemical Technology: Prague; FIZ CHEMIE GmbH: Berlin, 1993.
- 93-mal/woo Malhotra, R.; Woolf, L. A. Volumetric Measurements under Pressure for 2,2,4-Trimethylpentane at Temperatures up to 353.15 K and for Benzene and Three of their Mixtures at Temperatures up to 348.15 K. *Int. J. Thermophys.* **1993**, *14*, 1153–1172.
- 93-yok/ebi Yokoyama, C.; Ebina, T.; Takahashi, S. Melting Temperatures of Several Polycyclic Aromatic Compounds under High Pressures. *Fluid Phase Equilib.* **1993**, *84*, 207–223.
- 94-ami/ara Aminabhavi, T. M.; Aralaguppi, M. I.; Gopalakrishna, B.; Khinnavar, R. S. Densities, Shear Viscosities, Refractive Indices, and Speed of Sound of Bis(2-methoxyethyl) Ether with Hexane, Heptane, Octane, and 2,2,4-Trimethylpentane in the Temperature Interval 298.15–318.15 K. *J. Chem. Eng. Data* **1994**, *39*, 522–528.

- 94-pad/far Padua, A. A. H.; Fareleira, J. M. N. A.; Calado, J. C. G.; Wakeham, W. A. A Vibrating-wire Densimeter for Liquids at High Pressures: The Density of 2,2,4-Trimethylpentane from 298.15 to 348.15 K and up to 100 MPa. *Int. J. Thermophys.* **1994**, *15*, 229–243.
- 95-ami/gop Aminabhavi, T. M.; Gopalakrishna, B. J. Density, Viscosity, Refractive Index, and Speed of Sound in Binary Mixtures of 2-Ethoxyethanol with *n*-Alkanes ( $C_6$  to  $C_{12}$ ), 2,2,4-Trimethylpentane, and Cyclohexane in the Temperature Interval 298.15–313.15 K. *J. Chem. Eng. Data* **1995**, *40*, 632–641.
- 95-fuj/tam Fujii, S.; Tamura, K.; Murakami, S. Thermodynamic Properties of (an Alkylbenzene + Cyclohexane) at the Temperature 298.15 K. *J. Chem. Thermodyn.* **1995**, *27*, 1319–1328.
- 95-mie/osw Mier, W.; Oswald, G.; Tusel-Langer, E.; Lichtenhaler, R. N. Excess Enthalpy  $H^E$  of Binary Mixtures Containing Alkanes, Ethanol and Ethyl-*tert*-Butyl Ether (ETBE). *Ber. Bunsen-Ges. Phys. Chem.* **1995**, *99*, 1123–1130.
- 95-osw/pat Oswal, S.; Patel, A. T. Speeds of Sound, Isentropic Compressibilities, and Excess Volumes of Binary Mixtures. 2. Mono-*n*-alkylamines with Cyclohexane and Benzene. *J. Chem. Eng. Data* **1995**, *40*, 194–198.
- 95-pen/jac Penoncello, S. G.; Jacobsen, R. T.; Goodwin, A. R. H. A Thermodynamic Property Formulation for Cyclohexane. *Int. J. Thermophys.* **1995**, *16*, 519–531.
- 95-trc TRC Tables 23-2-(5.12010)-d. C–H. *n*-Alkenes (Monoolefins),  $C_6$ . *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1995; pp 2630–2631.
- 95-trc-1 TRC Tables 23-2-(5.12020)-d. C–H. *n*-Alkenes (Monoolefins),  $C_7$ . *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1995; pp 2650–2651.
- 95-trc-2 TRC Tables 23-2-(5.22020)-d. C–H. Alkadienes,  $C_3$  to  $C_5$ . *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1995; pp 2750–2751.
- 95-trc-3 TRC Tables 23-2-(3.10300)-d. C–H. *n*-Alkylcyclopentanes,  $C_5$  to  $C_{21}$ . *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1995; pp 1990–1990.
- 96-cib/hne Cibulka, I.; Hnědkovský, L. Liquid Densities at Elevated Pressures of *n*-Alkanes from  $C_5$  to  $C_{16}$ : A Critical Evaluation of Experimental Data. *J. Chem. Eng. Data* **1996**, *41*, 657–668.
- 96-hah/ulc Hahn, G.; Ulcay, K.; Svejda, P.; Siddiqi, M. A. Isothermal Compressibilities of Binary Liquid Mixtures of 1,2-Dichloroethane and of *trans*- and *cis*-1,2-Dichloroethene + *n*-Alkanes or 2,2,4-Trimethylpentane in the Pressure Range (0.1 to 10 MPa) and at 293.15 K. *J. Chem. Eng. Data* **1996**, *41*, 319–323.
- 96-pad/far Padua, A. A. H.; Fareleira, J. M. N. A.; Calado, J. C. G.; Wakeham, W. A. Validation of an Accurate Vibrating-Wire Densimeter: Density and Viscosity of Liquids over Wide Ranges of Temperature and Pressure. *Int. J. Thermophys.* **1996**, *17*, 781–802.
- 96-pad/far-1 Padua, A. A. H.; Fareleira, J. M. N. A.; Calado, J. C. G.; Wakeham, W. A. Density and Viscosity Measurements of 2,2,4-Trimethylpentane (Isooctane) from 198 K to 348 K and up to 100 MPa. *J. Chem. Eng. Data* **1996**, *41*, 1488–1494.
- 96-trc TRC Tables 23-2-(1.20000)-d. C–H. Alkanes (Paraffins),  $C_1$  to  $C_5$ . *TRC Thermodynamic Tables—Hydrocarbons*; Thermodynamics Research Center, The Texas A&M University System: College Station, TX, 1996; pp 1350–1353.
- 96-zab/ruz Záborský, M.; Růžička, V.; Majer, V.; Domalski, E. S. Heat Capacities of Liquids. Review and Recommended Values. *J. Phys. Chem. Ref. Data*; Monograph No. 6; American Chemical Society: Washington, DC, 1996.
- 97-bay/bon Baylaucq, A.; Boned, C.; Dague, P.; Lagourette, B. Measurements of the Viscosity and Density of Three Hydrocarbons and the Three Associated Binary Mixtures Versus Pressure and Temperature. *Int. J. Thermophys.* **1997**, *18*, 3–23.
- 97-oho/tam Ohomuro, K.; Tamura, K.; Murakami, S. Excess Volumes and Isentropic Compressibilities of the Binary Mixtures of Cyclohexanone with Globular Species at  $T = 298.15$  K. *J. Chem. Thermodyn.* **1997**, *29*, 287–294.
- 99-cib/tak Cibulka, I.; Takagi, T.  $P$ – $\rho$ – $T$  Data of Liquids: Summarization and Evaluation. 5. Aromatic Hydrocarbons. *J. Chem. Eng. Data* **1999**, *44*, 411–429.

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